

- 23) In which of the following sequences of measured numbers do all members of the sequence contain three significant figures?
- a) 3.03 and 3.30 and 0.033
 - b) 78,000 and 0.00780 and 780
 - c) 30.0 and 0.300 and 30,100
 - d) more than one correct response
 - e) no correct response

ANS: C PTS: 1 DIF: Moderate

KEY: Chemistry | general chemistry | significant figures NOT: Section 2.4

- 24) Which of the following digits in the measurement 654,300 seconds is an estimated digit?
- a) the last digit
 - b) the next to last zero
 - c) the three
 - d) more than one correct response
 - e) no correct response

ANS: C PTS: 1 DIF: Moderate

KEY: Chemistry | general chemistry | significant figures NOT: Section 2.4

- 25) Which of the following statements concerning the measured number 0.3030 is correct?
- a) Only one of the zeros in the number is significant.
 - b) Rounded off to two significant figures the number becomes 0.30.
 - c) Expressed in scientific notation the number becomes 3.03×10^{-1} .
 - d) More than one correct response.
 - e) No correct response.

ANS: B PTS: 1 DIF: Moderate

KEY: Chemistry | general chemistry | scientific notation; significant figures

NOT: Section 2.5

- 26) Which of the following mathematical expressions is correctly evaluated?
- a) $\frac{10^3}{10^{-4}} = 10^7$
 - b) $10^3 \times 10^4 = 10^{12}$
 - c) $\frac{10^3}{10^4} = 10^{-7}$
 - d) more than one correct response
 - e) no correct response

ANS: A PTS: 1 DIF: Moderate

KEY: Chemistry | general chemistry | scientific notation NOT: Section 2.6

- 27) Which of the following measured numbers contains three significant figures and has a magnitude of less than one?
- a) 3.30×10^5
 - b) 3.00×10^{-3}
 - c) 3.20×10^{-4}
 - d) more than one correct response
 - e) no correct response

ANS: D PTS: 1 DIF: Moderate
KEY: Chemistry | general chemistry | scientific notation NOT: Section 2.4

- 28) When expressed in scientific notation, the measured numbers 3200 and 3200.0 become, respectively,
- a) 3.2×10^3 and 3.200×10^3
 - b) 3.2×10^3 and 3.2000×10^3
 - c) 3.200×10^3 and 3.2000×10^3
 - d) more than one correct response
 - e) no correct response

ANS: B PTS: 1 DIF: Moderate
KEY: Chemistry | general chemistry | scientific notation NOT: Section 2.6

- 29) Which of the following measured numbers has an uncertainty of 0.01 associated with it?
- a) 32.930
 - b) 3.02×10^6
 - c) 3.0×10^{-1}
 - d) more than one correct response
 - e) no correct response

ANS: C PTS: 1 DIF: Easy
KEY: Chemistry | general chemistry | significant figures NOT: Section 2.4

- 30) Which of the following statements concerning conversion factors is incorrect?
- a) English-to-English conversion factors come from defined relationships
 - b) Metric-to-metric conversions come from measured relationships
 - c) English-to-English conversion factors always contain exact numbers
 - d) more than one correct response
 - e) no correct response

ANS: C PTS: 1 DIF: Moderate
KEY: Chemistry | general chemistry | conversion factors NOT: Section 2.4

31) Which of the following conversion factors would limit a calculation to two significant figures?

a) $\frac{453.6 \text{ g}}{1 \text{ lb}}$

b) $\frac{1 \text{ in.}}{2.54 \text{ cm}}$

c) $\frac{24 \text{ hr}}{1 \text{ day}}$

d) more than one correct response

e) no correct response

ANS: E PTS: 1 DIF: Moderate

KEY: Chemistry | general chemistry | conversion factors

NOT: Section 2.7

32) Density can be used as a conversion factor to convert from

a) mass to volume

b) volume to mass

c) metric unit mass to English unit mass

d) more than one correct response

e) no correct response

ANS: D PTS: 1 DIF: Moderate

KEY: Chemistry | general chemistry | density

NOT: Section 2.9

33) The density of table sugar is 1.59 g/mL. It is true that

a) 2.00 g of table sugar occupies a volume of 1.17 mL.

b) 3.00 g of table sugar occupies a volume of 1.97 mL.

c) 5.00 g of table sugar occupies a volume of 3.14 mL.

d) More than one correct response

e) No correct response

ANS: C PTS: 1 DIF: Difficult

KEY: Chemistry | general chemistry | density

NOT: Section 2.9

34) Which of the following statements concerning the three major temperature scales is *correct*?

a) Kelvin temperatures are always positive.

b) The equation for converting from Celsius to Kelvin involves the number 273.

c) The freezing point of water has a lower numerical value on the Kelvin scale than on the Fahrenheit scale.

d) More than one correct response.

e) No correct response.

ANS: D PTS: 1 DIF: Moderate

KEY: Chemistry | general chemistry | temperature scales

NOT: Section 2.10

- 35) In which of the following pairs of temperature readings are the two members of the pair equivalent to each other?
- 32°F and 273 K
 - 0°C and 373 K
 - 0°C and 40°F
 - more than one correct response
 - no correct response

ANS: A PTS: 1 DIF: Moderate

KEY: Chemistry | general chemistry | temperature scales NOT: Section 2.10

- 36) Statements:

- The meaning of a metric system prefix is independent of the base unit it modifies.
- “Trailing zeros” at the end of a measured number are never significant.
- The answer to the problem $10^5/10^{-3}$ is 10^2 .

- All three statements are true.
- Two of the three statements are true.
- Only one of the statements is true.
- None of the statements is true.

ANS: C PTS: 1 DIF: Moderate

KEY: Chemistry | general chemistry | metric system; scientific notation; significant figures

NOT: Section 2.2| Section 2.4| Section 2.6

- 37) Statements:

- In outer space, an astronaut may be weightless but never massless.
- The metric system prefixes *milli* and *micro* differ in mathematical meaning by a factor of 1000.
- The addition of 273 to a Fahrenheit temperature reading will convert it to a Kelvin temperature reading.

- All three statements are true.
- Two of the three statements are true.
- Only one of the statements is true.
- None of the statements is true.

ANS: B PTS: 1 DIF: Moderate

KEY: Chemistry | general chemistry | metric system; temperature scales

NOT: Section 2.10| Section 2.2

- 38) Statements:

- The measured number 2.410×10^{-3} contains three significant figures.
- The specific heat of water is higher than that of most other substances.
- The equation $1 \text{ kg} = 10^6 \text{ mg}$ is a correct mathematical statement.

- a) All three statements are true.
- b) Two of the three statements are true.
- c) Only one of the statements is true.
- d) None of the statements is true.

ANS: B PTS: 1 DIF: Moderate

KEY: Chemistry | general chemistry | metric system; significant figures; specific heat

NOT: Section 2.10| Section 2.11| Section 2.4

39) Statements:

(1) The answer to the addition problem $3.21 + 32 + 3.22$ should have an uncertainty of hundredths.

(2) The measurement 653,899, when rounded to five significant figures, becomes 65,390.

(3) The higher the specific heat of a substance, the more its temperature will change when it absorbs a given amount of heat.

- a) All three statements are true.
- b) Two of the three statements are true.
- c) Only one of the statements is true.
- d) None of the statements is true.

ANS: D PTS: 1 DIF: Moderate

KEY: Chemistry | general chemistry | significant figures; specific heat

NOT: Section 2.11| Section 2.5

40) Statements:

(1) The conversion factor 10^3 m/1 km contains an unlimited number of significant figures.

(2) Density may be used as a conversion factor to convert from mass to volume.

(3) The equation $2.33 \text{ lb} = 625 \text{ g}$ is a correct mathematical statement.

- a) All three statements are true.
- b) Two of the three statements are true.
- c) Only one of the statements is true.
- d) None of the statements is true.

ANS: B PTS: 1 DIF: Easy

KEY: Chemistry | general chemistry | conversion factors; density

NOT: Section 2.7| Section 2.9

41) Statements:

(1) A deciliter is equal to 100 milliliters.

(2) The Kelvin temperature scale is closely related mathematically to the Celsius temperature scale.

(3) Measurements cannot be exact because two estimated digits are always recorded as part of any measurement.

- a) All three statements are true.
- b) Two of the three statements are true.
- c) Only one of the statements is true.
- d) None of the statements is true.

ANS: B PTS: 1 DIF: Moderate

KEY: Chemistry | general chemistry | metric system; significant figures; temperature scales

NOT: Section 2.10| Section 2.2| Section 2.4

42) Statements:

(1) The answer to the calculation $12.00 \times (6.00 \times 10^{23})$ should contain three significant figures.

(2) A meter is slightly larger than a yard, and a liter is slightly larger than a quart.

(3) The numbers 3.30×10^{-1} and 3.30×10^1 both have a magnitude of less than one.

- a) All three statements are true.
- b) Two of the three statements are true.
- c) Only one of the statements is true.
- d) None of the statements is true.

ANS: B PTS: 1 DIF: Moderate

KEY: Chemistry | general chemistry | metric system; scientific notation; significant figures
| mathematical operations

NOT: Section 2.2| Section 2.5| Section 2.6

43) Statements:

(1) The size of the degree is the same on the Fahrenheit and Celsius temperature scales.

(2) The measurement 62,300 has an uncertainty of ± 100 .

(3) The answer to the calculation $8.45 + 10.40$ should contain four significant figures.

- a) All three statements are true.
- b) Two of the three statements are true.
- c) Only one of the statements is true.
- d) None of the statements is true.

ANS: B PTS: 1 DIF: Difficult

KEY: Chemistry | general chemistry | significant figures; temperature scales | mathematical
operations

NOT: Section 2.10| Section 2.4| Section 2.5

44) Statements:

(1) The measured numbers 244,000 and 0.000244 contain the same number of significant figures.

(2) One cubic centimeter is equal to ten milliliters.

(3) The conversion factor 1 in/2.54 cm, when used as written, would decrease unit size.

- a) All three statements are true.
- b) Two of the three statements are true.
- c) Only one of the statements is true.
- d) None of the statements is true.

ANS: C PTS: 1 DIF: Moderate

KEY: Chemistry | general chemistry | conversion factors; metric system; significant figures

NOT: Section 2.2| Section 2.4| Section 2.7