SPE 4th Edition Solution Manual Chapter 2.

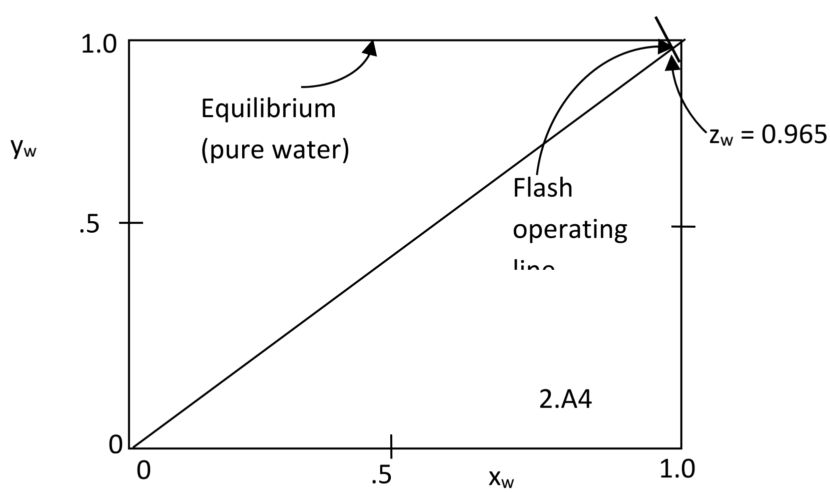
New Problems and new solutions are listed as new immediately after the solution number. These new problems are:2A8, 2A10 parts c-e, 2A11,2A12, 2A13, 2A14, 2C4, 2D1-part g, 2D3, 2D6, 2D7, 2D11, 2D13, 2D14, 2D20, 2D22, 2D23, 2D31, 2D32, 2E3, 2F4, 2G2, 2G3, 2H1, 2H3, 2H4, 2H5 and 2H6.

2.A1. Feed to flash drum is a liquid at high pressure. At this pressure its enthalpy can be calculated as a liquid. eg. $h\left(T_{F,P_{high}}\right) = c_{p_{LIQ}}\left(T_F - T_{ref}\right)$. When pressure is dropped the mixture is above its bubble point and is a two-phase mixture (It "flashes"). In the flash mixture enthalpy is unchanged but temperature changes. Feed location cannot be found from T_F and z on the graph because equilibrium data is at a lower pressure on the graph used for this calculation.

2.A2. Yes.

2.A3. The liquid is superheated when the pressure drops, and the energy comes from the amount of superheat.

2.A4.



2.A6. In a flash drum separating a multicomponent mixture, raising the pressure will:
i. Decrease the drum diameter and decrease the relative volatilities. *Answer is i.*

2.A8. *New Problem in 4th ed.*
. a. At 100°C and a pressure of 200 kPa what is the K value of n-hexane? 0.29
b. As the pressure increases, the K value
a. increases, b. decreases, c. stays constant b
c. Within a homologous series such as light hydrocarbons as the molecular weight increases, the K value (at constant pressure and temperature)
a. increases, b. decreases, c. stays constant b
d. At what pressure does pure propane boil at a temperature of -30°C? 160 kPa

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2.A9. a. The answer is 3.5 to 3.6

b. The answer is 36ºC

c. *This part is new in* 4^{th} *ed.* 102ºC

2.A10. *Parts c, d, and e are new in* 4^{th} *ed.* *a.* 0.22; b. No; c. From y-x plot for Methanol x = 0.65, $y_M = 0.85$; thus, $y_W = 0.15$. d. $K_M = 0.579/0.2 = 2.895$, $K_W = (1 - 0.579)/(1 - 0.2) = 0.52625$. e. $\alpha_{M\text{-}W} = K_M/K_W = 2.895/0.52625 = 5.501$.

2.A11. *New problem in* 4^{th} *edition.* Because of the presence of air this is not a binary system. Also, it is not at equilibrium.

2.A12. *New problem in* 4^{th} *edition.* The entire system design includes extensive variables and intensive variables necessary to solve mass and energy balances. Gibbs phase rule refers only to the intensive variables needed to set equilibrium conditions.

2A13. *New problem in* 4^{th} *edition.* Although V is an extensive variable, V/F is an intensive variable and thus satisfies Gibbs phase rule.

2A14. *New problem in* 4^{th} *edition.* 1.0 kg/cm^2 = 0.980665 bar = 0.96784 atm.
Source: http://www.unit-conversion.info/pressure.html

2.B1. Must be sure you don't violate Gibbs phase rule for intensive variables in equilibrium. Examples:

F, z, T_{drum}, P_{drum}	F, T_F, z, p	F, h_F, z, p
F, z, y, P_{drum}	F, T_F, z, y	F, h_F, z, y
F, z, x, p_{drum}	F, T_F, z, x	etc.
F, z, y, p_{drum}	$F, T_F, z, T_{drum},\ p_{drum}$	
F, z, x, T_{drum}	F, T_F, y, p	
Drum dimensions, $z, F_{drum},\ p_{drum}$	F, T_F, y, T_{drum}	
Drum dimensions, z, y, p_{drum}	F, T_F, x, p	
etc.	F, T_F, x, T_{drum}	
	F, T_F, y, x	

2.B2. This is essentially the same problem (disguised) as problem 2-D1c and e but with an existing (larger) drum and a higher flow rate.

With y = 0.58, x = 0.20, and V/F = 0.25 which corresponds to 2-D1c.

$$\text{If } F = 1000\frac{\text{lb mole}}{\text{hr}},\ D = .98 \text{ and } L\ =\ 2.95 \text{ ft from Problem 2-D1e}.$$

Since D α $\sqrt{V}$ and for constant V/F, V α F, we have D α $\sqrt{F}$.

With F = 25,000:

$$\sqrt{F_{new}/F_{old}} = 5,\ D_{new}\ = 5\ D_{old}\ = 4.90, \text{ and } L_{new} = 3\ D_{new}\ = 14.7.$$

Existing drum is too small.

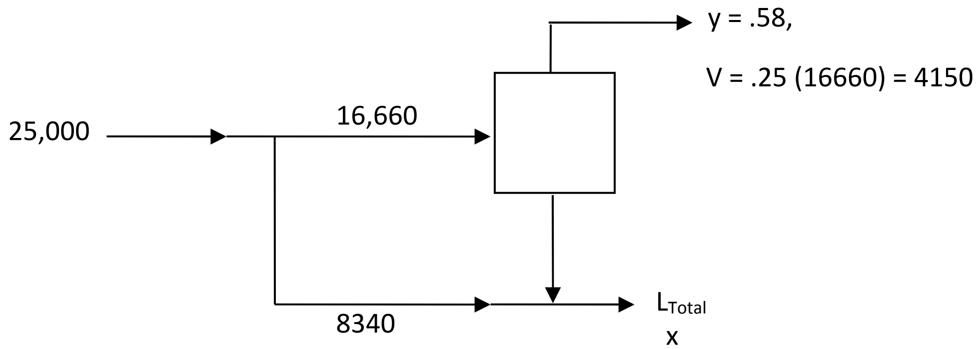
Feed rate drum can handle: F α D^2. $\frac{F_{existing}}{1000} = \left(\frac{D_{exist}}{.98}\right)^2 = \left(\frac{4}{.98}\right)^2$ gives

$F_{existing} = 16{,}660$ lbmol/h

Alternatives

a) Do drums in parallel. Add a second drum which can handle remaining 8340 lbmol/h.

b) Bypass with liquid mixing



Since x is not specified, use bypass. This produces less vapor.

c) Look at Eq. (2-62), which becomes

$$D = \sqrt{\frac{V(MW_v)}{3K_{drum}3600\sqrt{(\rho_L - \rho_v)\rho_v}}}$$

Bypass reduces V

c1) K_{drum} is already 0.35. Perhaps small improvements can be made with a better demister → Talk to the manufacturers.

c2) ρ_v can be increased by increasing pressure. Thus operate at higher pressure. Note this will change the equilibrium data and raise temperature. Thus a complete new calculation needs to be done.

d) Try bypass with vapor mixing.

e) Other alternatives are possible.

2.C2. $$\frac{V}{F} = \left[\frac{-z_A}{(K_B - 1)} - \frac{z_B}{(K_A - 1)}\right]$$

2.C5. a. Start with $x_i = \frac{Fz_i}{L + VK_i}$ and let $V = F - L$

$$x_i = \frac{Fz_i}{L + (F - L)K_i} \text{ or } x_i = \frac{z_i}{\frac{L}{F} + \left(1 - \frac{L}{F}\right)K_i}$$

Then $y_i = K_i x_i = \dfrac{K_i z_i}{\dfrac{L}{F} + \left(1 - \dfrac{L}{F}\right) K_i}$

From $\sum y_i - \sum x_i = 0$ we obtain $\sum \dfrac{(K_i - 1) z_i}{\dfrac{L}{F} + \left(1 - \dfrac{L}{F}\right) K_i} = 0$

2.C4. *New Problem.* Prove that the intersection of the operating and y = x lines for binary flash distillation occurs at the mole fraction of the feed.

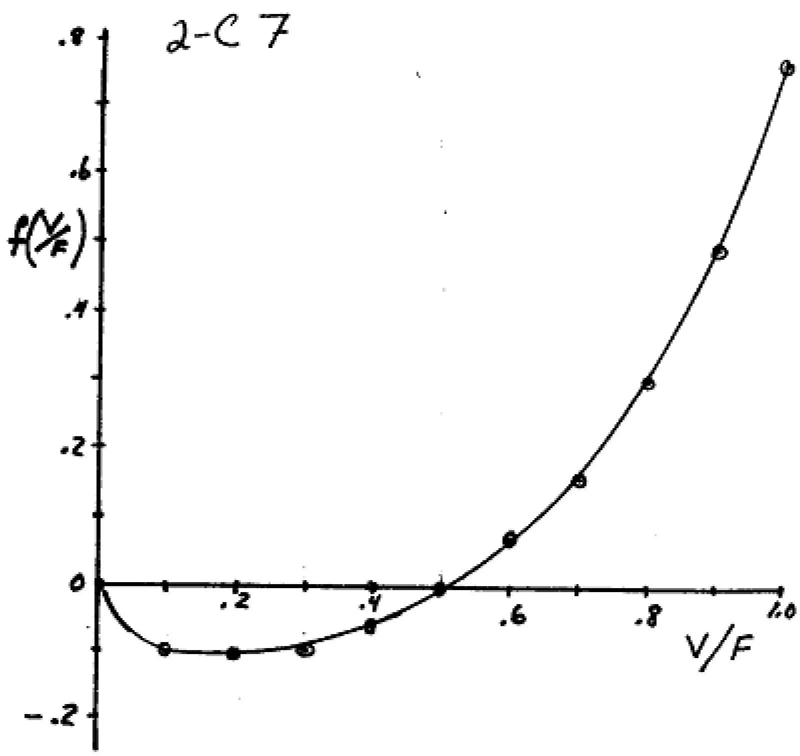
SOLUTION: $y = \frac{L}{V} y + \frac{F}{V} z$,rearrange: $y\left[1 + \frac{L}{V}\right] = \frac{F}{V} z$, or $y\left[\frac{V+L}{V}\right] = \frac{F}{V} z$ since V + L = F, the result is y = z and therefore

$$x = y = z \qquad \textbf{(2-18)}$$

The intersection is at the feed composition.

2.C7. $\sum \dfrac{z_i}{1 + (K_i - 1)\dfrac{V}{F}} - 1 = f\left(\dfrac{V}{F}\right)$ From data in Example 2-2 obtain:

V/F	0	.1	.2	.3	.4	.5	.6	.7	.8	.9	1.0
f	0	-.09	-.1	-.09	-.06	-.007	.07	.16	.3	.49	.77



2.C8. Derivation of Eqs. (2-62) and (2-63). Overall and component mass balances are,

$F = V + L_1 + L_2$ and $Fz_i = L_1 x_{i,L1} + L_2 + x_{i,L2} + Vy_i$ Substituting in Eqs. (2-60b) and 2-60c)

$$Fz_i = L_1 K_{i,L1-L2} x_{i,L2} + L_2 x_{i,L2} + VK_{iV-L2} x_{i,L2}$$

Solving,

$$x_{i,L2} = \frac{Fz_i}{L_1 K_{i,L2} + L_2 + VK_{i,V-L2}} = \frac{Fz_i}{L_1 K_{i,L1-L2} + F - V - L_1 + VK_{i,V-L2}}$$

Dividing numerator and denominator by F and collecting terms.

$$x_{i,liq2} = \frac{z_i}{1 + \left(K_{i,L1-L2} - 1\right)\frac{L_1}{F} + \left(K_{i,V-L2} - 1\right)\frac{V}{F}}$$

Since $y_i = K_{i,V-L2} x_{i,L2}$, $y_i = \dfrac{K_{i,V-L2} z_i}{1 + \left(K_{i,L1-L2} - 1\right)\frac{L_1}{F} + \left(K_{i,V-L2} - 1\right)\frac{V}{F}}$

Stoichiometric equations, $\sum_{i=1}^{C} x_{i,L2} = 1$, $\sum_{i=1}^{C} y_i = 1$, thus, $\sum_{i=1}^{C} y_i - \sum_{i=1}^{C} x_{i,L2} = 0$

which becomes

$$\sum_{i=1}^{C} \frac{\left(K_{i,V-L2} - 1\right) z_i}{\left[1 + \left(K_{i,L1-L2} - 1\right)\frac{L_1}{F} + \left(K_{i,V-L2} - 1\right)\frac{V}{F}\right]} = 0 \qquad (2\text{-}62)$$

Since $x_{i,liq1} = K_{i,L1-L2} x_{i,liq2}$, we have $x_{i,liq1} = \dfrac{K_{i,L1-L2} z_i}{1 + \left(K_{i,L1-L2} - 1\right)\frac{L_1}{F} + \left(K_{i,V-L2} - 1\right)\frac{V}{F}}$

In addition,

$$\sum x_{i,liq1} - \sum x_{i,liq2} = 0 = \sum_{i=1}^{C} \frac{\left(K_{i,L1-L2} - 1\right) z_i}{\left[1 + \left(K_{i,L1-L2} - 1\right)\frac{L_1}{F} + \left(K_{i,V-L2} - 1\right)\frac{V}{F}\right]} \qquad (2\text{-}63)$$

2.D1.

a. $V = (0.4)100 = 40$ and $L = F - V = 60$ kmol/h
Slope op. line $= -L/V = -3/2$, $y = x = z = 0.6$
See graph. $y = 0.77$ and $x = 0.48$

b. $V = (0.4)(1500) = 600$ and $L = 900$. Rest same as part a.

c. Plot $x = 0.2$ on equil. Diagram and $y = x = z = 0.3$. $y_{intercept} = zF/V = 1.2$
$V/F = z/1.2 = 0.25$. From equil $y = 0.58$.

d. Plot $x = 0.45$ on equilibrium curve.

$$\text{Slope} = -\frac{L}{V} = -\frac{F-V}{V} = -\frac{1-V/F}{V/F} = \frac{-.8}{.2} = -4$$

Plot operating line, $y = x = z$ at $z = 0.51$. From mass balance $F = 37.5$ kmol/h.

e. *Find Liquid Density.*

$$\overline{MW}_L = x_m (MW_m) + x_w (MW_w) = (.2)(32.04) + (.8)(18.01) = 20.82$$

Then, $\bar{V}_L = x_m \dfrac{MW_m}{\rho_m} + x_w \dfrac{MW_w}{\rho_w} = .2\left(\dfrac{32.04}{.7914}\right) + .8\left(\dfrac{18.01}{1.00}\right) = 22.51$ ml/mol

$$\rho_L = \overline{MW}_L / \bar{V}_L = 20.82/22.51 = 0.925 \text{ g/ml}$$

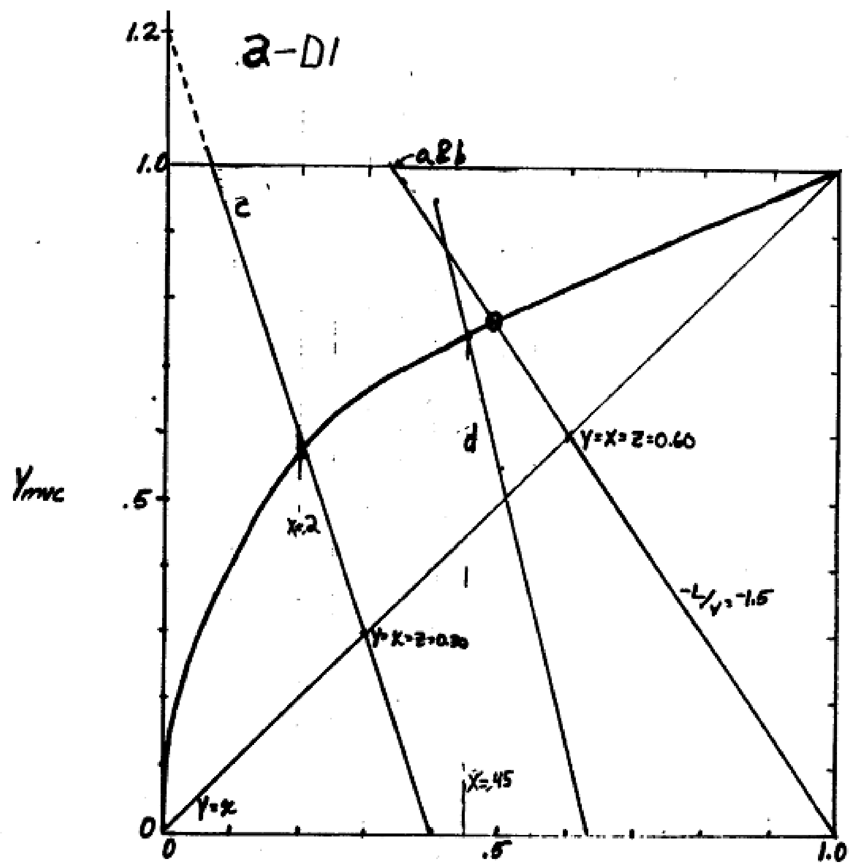
Vapor Density: $\rho_V = p(MW)_{V,avg}/RT$ (Need temperature of the drum)

$$\overline{MW}_v = y_m (MW)_m + y_w (MW)_w = .58(32.04) + .42(18.01) = 26.15 \text{ g/mol}$$

Find Temperature of the Drum T: From Table 3-3 find T when

$y = .58,\ x = 20,\ T = 81.7°C = 354.7K$

$$\rho_v = (1 \text{ atm})(26.15 \text{ g/mol}) \Big/ \left[\left(82.0575 \frac{\text{ml atm}}{\text{mol °K}}\right)(354.7 \text{ K})\right] = 8.98 \times 10^{-4} \text{ g/ml}$$



Find Permissible velocity:

$$u_{perm} = K_{drum}\sqrt{(\rho_L - \rho_v)/\rho_v},\ K_{drum} = \exp\left[A + B(\ell nF_{lv}) + C(\ell nF_{lv})^2 + D(\ell nF_{lv})^3 + E(\ell nF_{lv})^4\right] \qquad 2\text{-}60$$

$$V = \left(\frac{V}{F}\right)F = (0.25)1000 = 250 \text{ lbmol/h},\ W_v = V\left(\overline{MW}_v\right) = 250\left(26.15\frac{\text{lb}}{\text{lbmol}}\right) = 6537.5 \text{ lb/h}$$

$$L = F - V = 1000 - 250 = 750 \text{ lbmol/h, and } W_L = (L)\left(\overline{MW}_L\right) = (750)(20.82) = 15{,}615 \text{ lb/h,}$$

$$F_{lv} = \frac{W_L}{W_V}\sqrt{\frac{\rho_V}{\rho_L}} = \left(\frac{15615}{6537.5}\right)\sqrt{\frac{8.89\times10^{-4}}{.925}} = 0.0744, \text{ and } \ell n\left(F_{lv}\right) = -2.598$$

Then $K_{drum} = .442$, and $u_{perm} = .442\sqrt{\frac{.925 - 8.98\times10^{-4}}{8.98\times10^{-4}}} = 14.19$ ft/s

$$A_{cs} = \frac{V\left(\overline{MW}_v\right)}{u_{perm}3600\rho_v} = \frac{250(26.15)(454 \text{ g/lb})}{(14.19)(3600)\left(8.98\times10^{-4} \text{ g/ml}\right)\left(28316.85 \text{ ml/ft}^3\right)} = 2.28 \text{ ft}^2.$$

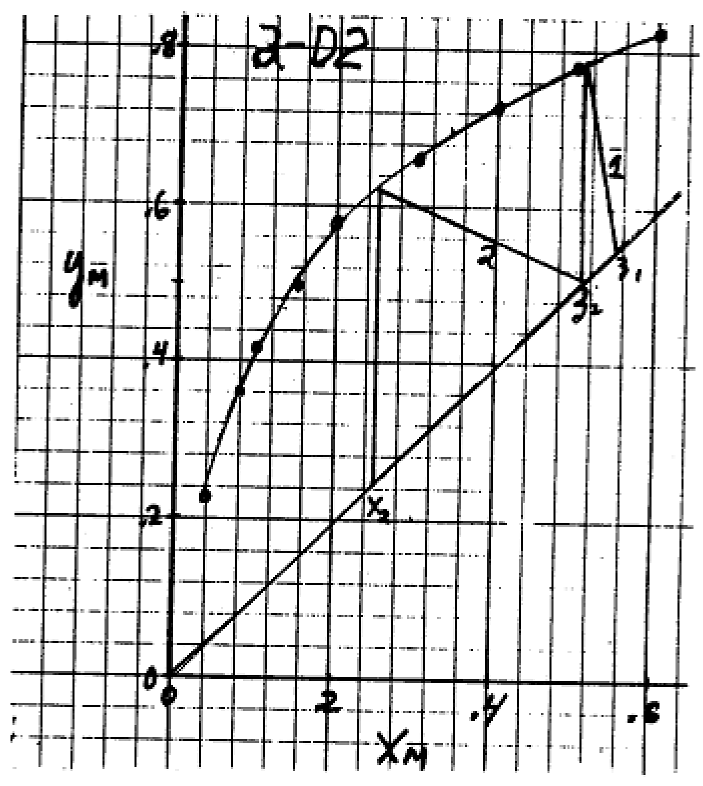
$D = \sqrt{4A_{cs}/\pi} = 1.705$ ft. Use 2 ft diameter. L ranges from $3\times D = 6$ ft to $5\times$ D=10 ft

Note that this design is conservative if a demister is used.

f. Plot T vs x from Table 3-3. When $T = 77^\circ C$, $x = 0.34$, $y = 0.69$. This problem is now very similar to 3-D1c. Can calculate V/F from mass balance, $Fz = Lx + Vy$. This is $Fz = (F - V)x + Vy$ or $\frac{V}{F} = \frac{z-y}{y-x} = \frac{0.4 - 0.34}{0.69 - 0.34} = 0.17$

g. *Part g is a new problem.* V = 16.18 mol/h, L = 33.82, y= 0.892, x = 0.756.

2-D2. Work backwards. Starting with x_2, find y_2 = 0.62 from equilibrium. From equilibrium point plot op. line of slope $= -(L/V)_2 = -\left(1 - \frac{V}{F}\right)_2 \Big/ (V/F)_2 = -3/7$. Find $z_2 = 0.51 = x_1$ (see Figure). From equilibrium, $y_1 = 0.78$ .For stage 1, $\frac{V}{F} = \frac{z_1 - x_1}{y_1 - x_1} = \frac{0.55 - 0.51}{0.78 - 0.51} = 0.148$.



2.D3. *New Problem in 4th edition..* Part a.

x ethane	T °C	y ethane
0	63.19	0
.025	56.18	0.1610
.05	49.57	0.2970
.10	37.57	0.5060
.15	27.17	0.6503
.20	18.26	0.7492
.25	10.64	0.8175
.30	4.11	0.8652
1.0	-37.47	1.0

b. See Figure. a. If 1 bubble of vapor product (V/F = 0) vapor product, vapor $y_E = 0.7492$ (highest) liquid $x_E = z_E = 0.20$ (highest) and T = 18.26 °C. If 1 drop of liquid product (V/F =1) $y_E = z_E = 0.20$ (lowest), $x_E = 0.035$, T (by linear interpolation) ~ 56.18 + [(49.57 – 56.18)/(.297 - .161)][.2 – 0.16] = 54.2 °C (highest).
c. See figure. Slope = -L/V = - (1 – V/F)/(V/F) = - .6/.4 = - 1.5. $x_E = 0.12$, $y_E = 0.57$, T = 33.4°C.
d. From equilibrium data $y_E = 0.7492$. For an F = 1, L = 1 – V, Ethane balance: .2L = 1(.3) – 0.7492 V. Solve 2 equations: V/F = 0.1821. Can also find V/F from slope of operating line.
e. If do linear interpolation on equilibrium data, x = 0.05 +(45-49.57)(0.1 -0.05)/(37.57 – 49.57) = 0.069. From equilibrium plot y = 0.375.
Mass balance for basis F = 1, L = 1 – V and 0.069 L = 0.18 – 0.375 V. Solve simultaneously, V/F = 0.363.

2.D4. *New problem in 3rd edition.* Highest temperature is dew point $(V/F = 0)$

Set $z_i = y_i$. $K_i = y_i/x_i$. Want $\sum x_i = \sum y_i/K_i = 1.0$

$$K_{ref}(T_{New}) = K_{ref}(T_{Old}) \quad \left(\sum(y_i/K_i)\right)$$

If pick C4 as reference: First guess $K_{butane} = 1.0$, $T = 41^\circ C$: $K_{C3} = 3.1, K_{C6} = 0.125$

$$\sum \frac{y_i}{K_i} = \frac{.2}{3.1} + \frac{.35}{1.0} + \frac{.45}{.125} = 4.0145 \quad \text{T too low}$$

Guess for reference: $K_{C4} = 4.014$, $T = 118^\circ C$: $K_{C3} = 8.8$, $K_{C6} = .9$

$$\sum \frac{y_i}{K_i} = \frac{.2}{8.8} + \frac{.35}{4.0145} + \frac{.45}{.9} = 0.6099$$

$$K_{C4,NEW} = 4.0145(.6099) = 2.45, T = 85: K_{C2} = 6.0, K_{C6} = 0.44$$

$$\sum \frac{y_i}{K_i} = \frac{.2}{6} + \frac{.35}{2.45} + \frac{.45}{.44} = 1.20$$

$$K_{C4,NEW} = 2.45 \times 1.2 = 2.94,\ T = 96^\circ C:\ K_{C3} = 6.9,\ K_{C6} = 0.56$$

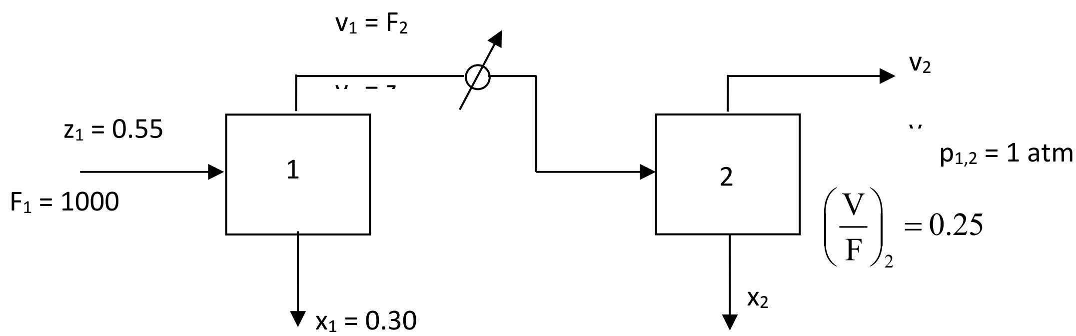
$$\sum \frac{y_i}{K_i} = \frac{.2}{6.9} + \frac{.35}{2.94} + \frac{.45}{.56} = 0.804 \Rightarrow \text{Gives } 84^\circ C$$

Use 90.5° → Avg last two T $\quad K_{C4} = 2.7,\ \ K_{C3} = 6.5,\ \ K_{C6} = 0.49$

$$\sum (y_i/K_i) = \frac{.2}{6.5} + \frac{.35}{2.7} + \frac{.45}{.49} = 1.079,\ T \sim 87-88^\circ C$$

Note: hexane probably better choice as reference.

2.D5. a)



b) $y_1 = -\frac{L}{V_1} x_1 + \frac{F}{V_1} z$ Plot 1st Op line. $\boxed{y_1 = 0.66 = z_2}$

$y = x = z = 0.55$ to $x_1 = 0.3$ on eq. curve (see graph)

$$\text{Slope } = -\frac{L}{V_1} = \frac{0.55 - 0.80}{.55 - 0} = -\frac{.25}{.55} = -0.454545 \quad L_1 + V_1 = F_1 = 1000$$

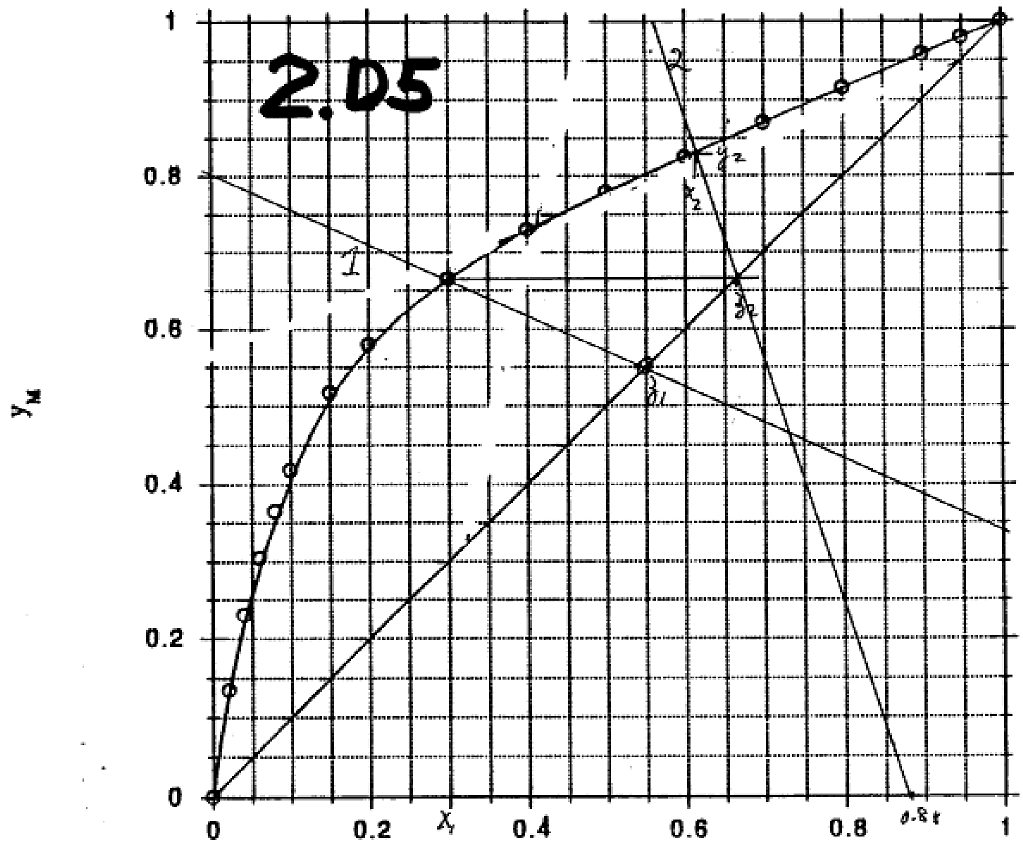
$\boxed{V_1 = 687.5 \text{ kmol/h} = F_2}$

$$\boxed{\left.\frac{V}{F}\right)_1 = \frac{687.5}{1000} = 0.6875}$$

c) Stage 2 $= \frac{V}{F} = 0.25$, $\quad -\frac{L}{V} = \frac{-0.75F}{0.25F} = -3$, $\quad y = x = z_2 = 0.66.$ Plot op line

At $x = 0,\ y = z/(V/F) = \frac{0.66}{0.25} = 2.64.$ At $y = 0,\ x_2 = \frac{F}{L} z = \frac{z}{L/F} = \frac{0.66}{0.75} = 0.88$

From graph $\boxed{y_2 = 0.82,\ x_2 = 0.63}$. $V_2 = \left(\frac{V}{F}\right)_2 F_2 = (0.25)687.5 = \boxed{171.875 \text{ kmol/h}}$



2.D6. *New problem in 4th ed.* a.) The answer is VP = 19.30 mm Hg

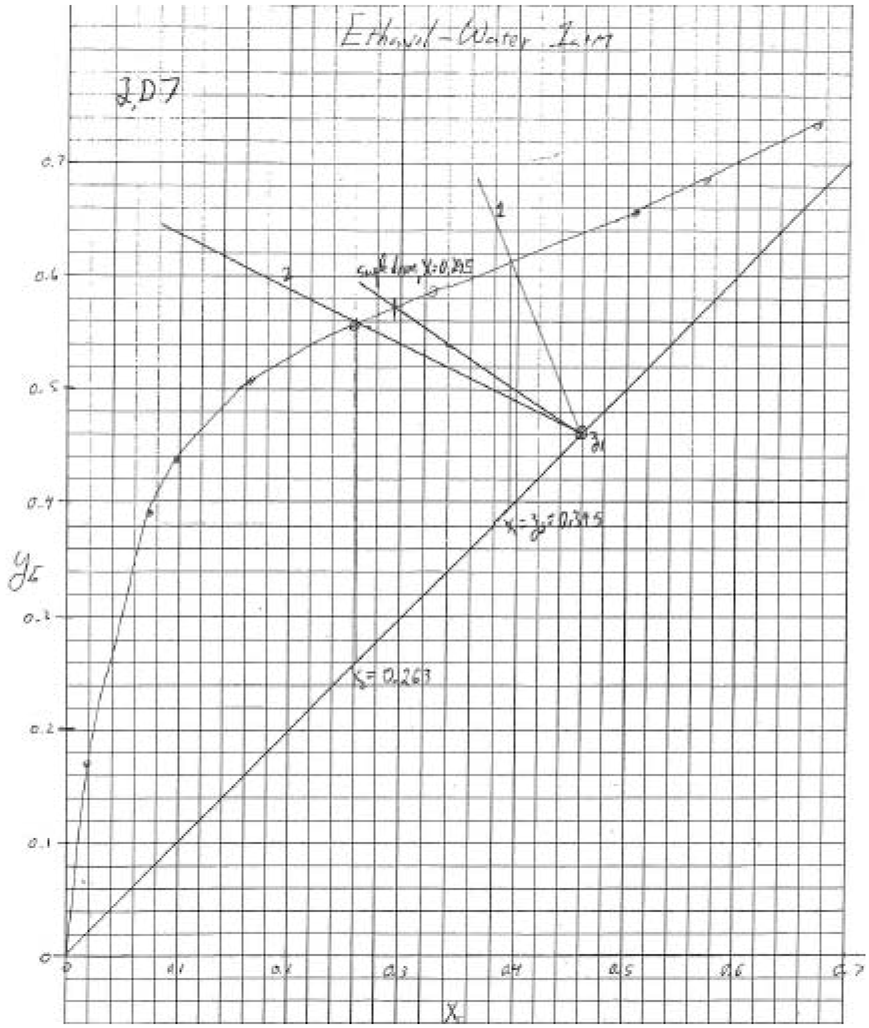
$$\log_{10}(\mathrm{VP}) = 6.8379 - \frac{1310.62}{100 + 136.05} = 1.2856$$

b.) The answer is K = 0.01693. $\mathrm{K} = \frac{\mathrm{VP}}{\mathrm{P_{tot}}} = \frac{19.30}{1.5(760)}$

2.D7. *New problem 4th ed.*

Part a. Drum 1: $V_1/F_1 = 0.3$, Slope op line = $-L/V = -.7/.3 = -7/3$, $y=x=z_1 = 0.46$. $L_1 = F_2 = 70$.
From graph $x_1 = z_2 = 0.395$
Drum 2: $V_1/F_1 = 30/70$, Slope op line = $-L/V = -7/3$, $y=x=z_2 = 0.395$. $L_1 = F_2 - V_2 = 40$.
From graph $x_2 = 0.263$

Part b. Single drum: $V/F = 0.6$, Slope op line = $-L/V = -40/60 = -2/3$, From graph $x = 0.295$.
More separation with 2 drums.



2.D8. Use Rachford-Rice eqn: $f\left(\frac{V}{F}\right) = \sum \frac{(K_i - 1) z_i}{1 + (K_i - 1) V/F} = 0$. Note that 2 atm = 203 kPa.

Find K_i from DePriester Chart: $K_1 = 73,\ K_2 = 4.1\ K_3 = .115$

Converge on $V/F = .076,\ V = F(V/F) = 152$ kmol/h, $L = F - V = 1848$ kmol/h .

From $x_i = \frac{z_i}{1 + \frac{V}{F}(K_i - 1)}$ we obtain $x_1 = .0077,\ x_2 = .0809,\ x_3 = .9113$

From $y_i = K_i x_i$, we obtain $y_1 = .5621,\ y_2 = .3649,\ y_3 = .1048$

2.D9. Need h_F to plot on diagram. Since pressure is high, feed remains a liquid

$$h_F = \bar{C}_{P_L}(T_F - T_{ref}),\quad T_{ref} = 0^\circ \text{ from chart}$$

$$\bar{C}_{P_L} = C_{P_{EtOH}} x_{EtOH} + C_{P_w} x_w$$

Where x_{EtOH} and x_w are *mole* fractions. Convert weight to mole fractions.

Basis: 100 kg mixture: 30 kg EtOH $= \frac{30}{46.07} = 0.651$ kmol

70 kg water $= 70/18.016 = 3.885$ Total = 4.536 kmol

Avg. MW $= \dfrac{100}{4.536} = 22.046$ Mole fracs: $x_E = \dfrac{0.6512}{4.536} = 0.1435,\ x_w = 0.8565$.

Use $C_{P_{L_{EtOH}}}$ at 100°C as an average C_P value.

$$\bar{C}_{P_L} = 37.96(.1435) + 18.0(.8565) = 20.86 \frac{\text{kcal}}{\text{kmol }^\circ\text{C}}$$

Per kg this is $\dfrac{\bar{C}_{P_L}}{MW_{avg}} = \dfrac{20.86}{22.046} = 0.946 \dfrac{\text{kcal}}{\text{kg }^\circ\text{C}}$

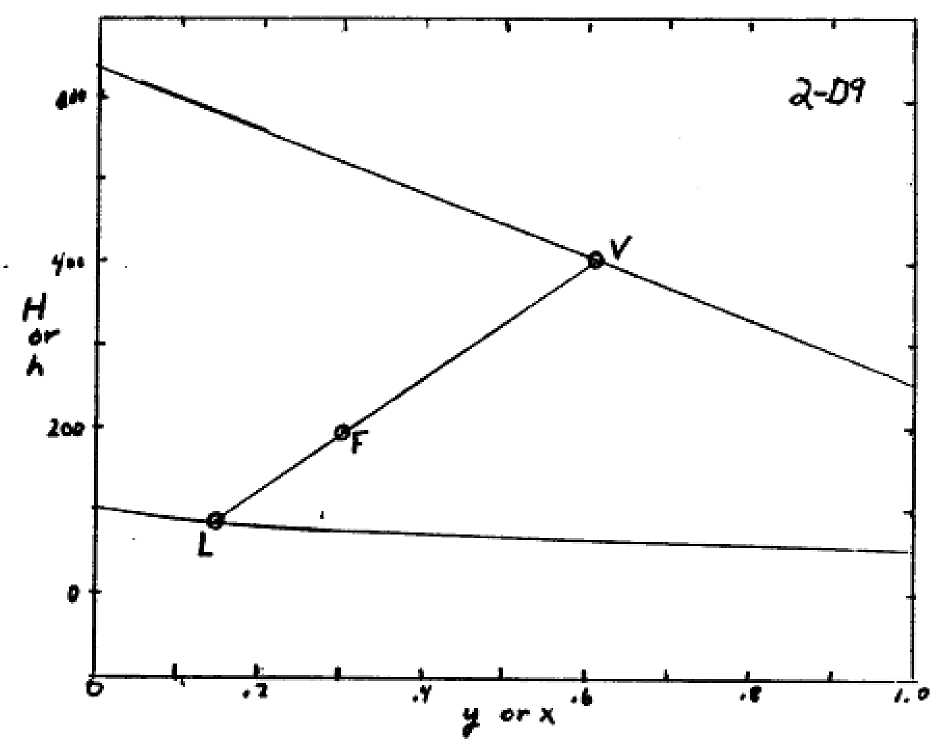
$$h_F = 0.946(2000) = 189.2 \text{ kcal/kg}$$

which can now be plotted on the enthalpy composition diagram.

Obtain $T_{drum} \approx 88.2\ ^\circ\text{C},\ x_E = 0.146,$ and $y_E = 0.617$.

For $F = 1000$ find L and V from F = L + V and Fz = Lx + Vy

which gives V = 326.9, and L = 673.1



Note: If use wt. fracs. $\bar{C}_{P_L} = 23.99$ & $\bar{C}_{P_L}/MW_{avg} = 1.088$ and $h_F = 217.6$. All *wrong*.

2.D.10 *Solution* 400 kPa, 70ºC $\quad z_{C4} = 35$ Mole % n-butane $\quad x_{C6} = 0.7$

From DePriester chart $\quad K_{C3} = 5,\quad K_{C4} = 1.9,\quad K_{C6} = 0.3$

Know $y_i = K_i x_i,\quad x_i = \dfrac{z_i}{1 + (K_i - 1)\dfrac{V}{F}},\quad \sum x_i = \sum y_i = 1 = \sum z_i$

R.R. $\quad \sum \frac{(K_i - 1) z_i}{1 + (K_i - 1)\frac{V}{F}} = 0 \qquad z_{C3} = 1 - z_{C6} - z_{C4} = .65 - z_{C6}$

C6: $0.7 = \frac{z_{C6}}{1 + (K_{C6} - 1)\frac{V}{F}} = \frac{z_{C6}}{1 - 0.7\frac{V}{F}} \Rightarrow z_{C6} = 0.7\left(1 - 0.7\frac{V}{F}\right),\ z_{C6} = 0.7 - 0.49\frac{V}{F}$

RR Eq: $\quad \frac{4(.65 - z_{C6})}{1 + 4\frac{V}{F}} + \frac{0.9(.35)}{1 + 0.9\frac{V}{F}} - \frac{0.7 z_{C6}}{1 - 0.7\frac{V}{F}} = 0$

2 equations & 2 unknowns. Substitute in for z_{C6}. Do in Spreadsheet. Use Goal – Seek to find V/F. V/F = 0.594 when R.R. equation $= 0.000881$.

$z_{C6} = 0.7 - 0.49\frac{V}{F} = 0.7 - (0.49)(0.594) = 0.40894$

2.D11. *New Problem 4th ed.* Obtain K ethylene = 2.2, K propylene = 0.56 from De Priester chart.

$K_E = y_E/x_E$ and $K_P = y_P/x_P$ Since $y_p = 1 - y_E$ and $x_p = 1 - x_E$, $K_p = (1 - y_E)/(1 - x_E)$.

Thus, 2 eqs and 2 unknowns. Solve for y_E and x_E.

$x_E = (1 - K_p) / (K_E - K_p)$ and $y_E = K_E x_E = K_E (1 - K_p) / (K_E - K_p)$

$x_E = (1 - 0.56) / (2.2 - 0.56) = 0.268$ and $y_E = K_E x_E = (2.2)(0.268) = 0.590$

Check: $x_p = 1 - x_E = 1 - 0.268 = 0.732$ and $y_p = 1 - y_E = 1 - 0.590 = 0.410$

$K_p = y_p/x_p = 0.410/ 0.732 = 0.56$ OK

2.D12. For problem 2.D1c, plot x = 0.2 on equilibrium diagram with feed composition of 0.3. The resulting operating line has a y intercept $= z/(V/F) = 1.2$. Thus $V/F = 0.25$ (see figure in Solution to 2.D1) Vapor mole fraction is y = 0.58.

Find Liquid Density.

$$\overline{MW}_L = x_m (MW_m) + x_w (MW_w) = (.2)(32.04) + (.8)(18.01) = 20.82$$

Then, $\bar{V}_L = x_m \frac{MW_m}{\rho_m} + x_w \frac{MW_w}{\rho_w} = .2\left(\frac{32.04}{.7914}\right) + .8\left(\frac{18.01}{1.00}\right) = 22.51$ ml/mol

$$\rho_L = \overline{MW}_L / \bar{V}_L = 20.82/22.51 = 0.925 \text{ g/ml}$$

Find Vapor Density. $\rho_v = \frac{p\left(\overline{MW}\right)_v}{RT}$ (Need temperature of the drum)

$$\overline{MW}_v = y_m (MW)_m + y_w (MW)_w = .58(32.04) + .42(18.01) = 26.15 \text{ g/mol}$$

Find Temperature of the Drum T:

From Table 2-7 find T corresponding to $y = .58,\ x = 20,\ T = 81.7^\circ C = 354.7K$

$$\rho_v = (1 \text{ atm})(26.15 \text{ g/mol}) \Big/ \left[\left(82.0575 \frac{\text{ml atm}}{\text{mol }^\circ K}\right)(354.7 \text{ K})\right] = 8.98\ 10^{-4} \text{ g/ml}$$

Find Permissible velocity: $\quad u_{perm} = K_{drum}\sqrt{(\rho_L - \rho_v)/\rho_v}$

$$K_{drum,horizontal} = 1.25 \times K_{drum,vertical} = \left\{ \exp\left[A + B\left(\ell nF_{lv}\right) + C\left(\ell nF_{lv}\right)^2 + D\left(\ell nF_{lv}\right)^3 + E\left(\ell nF_{lv}\right)^4 \right] \right\} \times 1.25$$

Since $V = (V/F) = (0.25)1000 = 250$ lbmol/h,

$$W_v = V\left(\overline{MW}_v\right) = 250(26.15 \text{ lb/lbmol}) = 6537.5 \text{ lb/h}$$

$L = F - V = 1000 - 250 = 750$ lbmol/h, and $W_L = (L)\left(\overline{MW}_L\right) = (750)(20.82) = 15,615$ lb/h,

$$F_{lv} = \frac{W_L}{W_V}\sqrt{\frac{\rho_V}{\rho_L}} = \left(\frac{15615}{6537.5}\right)\sqrt{\frac{8.98\times10^{-4}}{.925}} = 0.0744, \text{ and } \ell n\left(F_{lv}\right) = -2.598$$

$$K_{drum,vertical} = 0.442, \text{ and } K_{drum,horiz} = 0.5525$$

$$u_{perm} = 0.5525\sqrt{\frac{0.925 - 8.98\times10^{-4}}{8.98\times10^{-4}}} = 17.74 \text{ ft/s}$$

$$A_{cs} = \frac{V\left(\overline{MW}_v\right)}{u_{perm}3600\rho_v} = \frac{250(26.15)(454 \text{ g/lbm})}{(17.74)(3600)\left(8.98\times10^{-4} \text{ g/ml}\right)\left(28316.85 \text{ ml/ft}^3\right)}$$

$$A_{Cs} = 1.824 \text{ ft}^2, \quad A_T = A_{Cs}/0.2 = 9.12 \text{ ft}^2$$

With L/D = 4, $\quad D = \sqrt{4A_T/\pi} = 3.41$ ft and $L = 13.6$ ft

2.D13. *New Problem 4th ed.* $x_{butane} = 1 - x_E = 0.912$, $y_{butane} = 1 - y_E = 0.454$. $K_E = y_E/x_E = 0.546/0.088 = 6.20$, $K_{butane} = y_B/x_B = 0.454/0.912 = 0.498$.

Plot K_E and K_{butane} on DePriester chart. Draw straight line between them. Intersections with T and P axis give $T_{drum} = 15$ °C, and $p_{drum} = 385$ kPa from Figure 2-12.

Use mass balances to find V/F: F = L + V and $Fz_E = Lx_E + Vy_E$. Substitute L = F – V into ethane balance and divide both sides by F. Obtain: z = (1 – V/F)x + y(V/F).

Solve for V/F = (z-x)/(y-x) = (0.36 – 0.088)/(0.546 – 0.088) = 0.594.

Spreadsheet used as a check (using T=15 and p = 385) gave V/F = 0.593.

2.D14. *New Problem 4th ed.* DePriester chart, Fig. 2-12: $K_{C1} = 50$, $K_{C4} = 1.1$, and $K_{C5} = 0.37$; $z_1 = 0.12$, $z_4 = 0.48$, $z_5 = 0.40$

Rachford-Rice equation: $$\frac{(K_{C2}-1)z_{C1}}{1+(K_{C1}-1)\frac{V}{F}} + \frac{(K_{iC4}-1)z_{nC4}}{1+(K_{nC4}-1)\frac{V}{F}} + \frac{(K_{nC4}-1)z_{nC5}}{1+(K_{nC5}-1)\frac{V}{F}} = 0$$

Equation becomes: $$\frac{5.88}{1+49(V/F)} + \frac{0.048}{1+0.1(V/F)} - \frac{0.252}{1-0.63(V/F)} = 0$$

Trials: V/F = 0.4, Eq. = -.005345; V/F = 0.39, Eq. = 0.004506; V/F = 0.394, Eq. = 0.000546, which is close enough with DePriester chart.

Liquid mole fractions:

$$x_{C1} = \frac{z_{C1}}{1-(K_{C1}-1)(V/F)} = \frac{.12}{1+49(.394)} = 0.00591;$$ $x_{C4} = 0.4618$, $x_{C5} = 0.5321$, and $\sum x_i = 0.9998$

Vapor mole fractions: $y_i = K_i x_i$: $y_{C1} = 50(0.00591) = 0.2955$, $y_{C4} = 0.5080$, $y_{C5} = 0.1969$, $\sum y_i = 1.0004$.

2.D15. This is an unusual way of stating problem. However, if we count specified variables we see that problem is not over or under specified. Usually V/F would be the variable, but here it isn't. We can still write R-R eqn. Will have three variables: z_{C2}, z_{iC4}, z_{nC4}. Need two other eqns: $z_{iC4}/z_{nC4} =$ constant, and $z_{C2} + z_{iC4} + z_{nC4} = 1.0$

Thus, solve three equations and three unknowns simultaneously.

Do It. Rachford-Rice equation is,

$$\frac{(K_{C2}-1)z_{C2}}{1+(K_{C2}-1)\frac{V}{F}} + \frac{(K_{iC4}-1)z_{iC4}}{1+(K_{iC4}-1)\frac{V}{F}} + \frac{(K_{nC4}-1)z_{nC2}}{1+(K_{nC4}-1)\frac{V}{F}} = 0$$

Can solve for $z_{C2} = 1 - z_{iC4}$ and $z_{iC4} = (.8)\, z_{nC4}$. Thus $z_{C2} = 1 - 1.8\, z_{nC4}$

Substitute for z_{iC4} and z_{C2} into R-R eqn.

$$\frac{(K_{C2}-1)}{1+(K_{C2}-1)\frac{V}{F}}(1-1.8\, z_{nC4}) + \frac{.8(K_{iC4}-1)}{1+(K_{iC4}-1)\frac{V}{F}} z_{nC4} + z_{nC4}\frac{(K_{nC4}-1)}{1+(K_{nC4}-1)\frac{V}{F}} = 0$$

Thus,

$$z_{nC4} = \frac{\dfrac{(K_{C2}-1)}{1+(K_{C2}-1)\frac{V}{F}}}{1.8\dfrac{(K_{C2}-1)}{1+(K_{C2}-1)\frac{V}{F}} - \dfrac{.8(K_{iC4}-1)}{1+(K_{iC4}-1)\frac{V}{F}} - \dfrac{(K_{nC4}-1)}{1+(K_{nC4}-1)\frac{V}{F}}}$$

Can now find K values and plug away. $K_{C2} = 2.92$, $K_{iC4} = .375$, $K_{nC4} = .26$.

Solution is $z_{nC4} = 0.2957$, $z_{iC4} = .8\ (.2957) = 0.2366$, and $z_{C2} = 0.4677$

2.D16. $z_{C1} = 0.5,\ z_{C4} = 0.1,\ z_{C5} = 0.15,\ z_{C6} = 0.25,\ K_{C1} = 50,\ K_{C4} = .6,\ K_{C5} = .17,\ K_{C6} = 0.05$

1st guess. Can assume all C_1 in vapor, ~ 1/3 C_4 in vapor, C_5 & C_6 in bottom

$$V/F)_1 = .5 + (.1)/3 = .53 \quad \text{This first guess is not critical.}$$

R.R. eq. $$f\left(\frac{V}{F}\right) = \sum \frac{(K_i-1)z_i}{1+(K_i-1)V/F} = 0$$

$$\frac{49(.5)}{1+49(.53)} + \frac{(-.4)(.1)}{1-.4(.53)} + \frac{(-.83)(.15)}{1-.83(.53)} + \frac{(-.95)(.25)}{1-.95(.53)} = 0.157$$

Eq. 3.33 $$\left.\frac{V}{F}\right)_2 = \left(\frac{V}{F}\right)_1 + \frac{f(V/F)_1}{\sum \dfrac{z_i(K_i-1)^2}{\left[1-(K_i-1)\frac{V}{F}\right]^2}}$$

where $(V/F)_1 = 0.53$ and $f(V/F)_1 = 0.157$.

calculate $(V/F)_2 = .53 + 0.157/2.92 = 0.584$

$$V = .584(150) = 87.6 \text{ kmol/h and } L = 150 - 87.6 = 62.4$$

$$x_{C1} = \frac{z_{C1}}{1-(K_{C1}-1)(V/F)} = \frac{.5}{1+49(.584)} = 0.016883$$

$$y_{C1} = K_{C1}\ x_{C1} = 50(0.016883) = 0.844$$

Similar for other components.

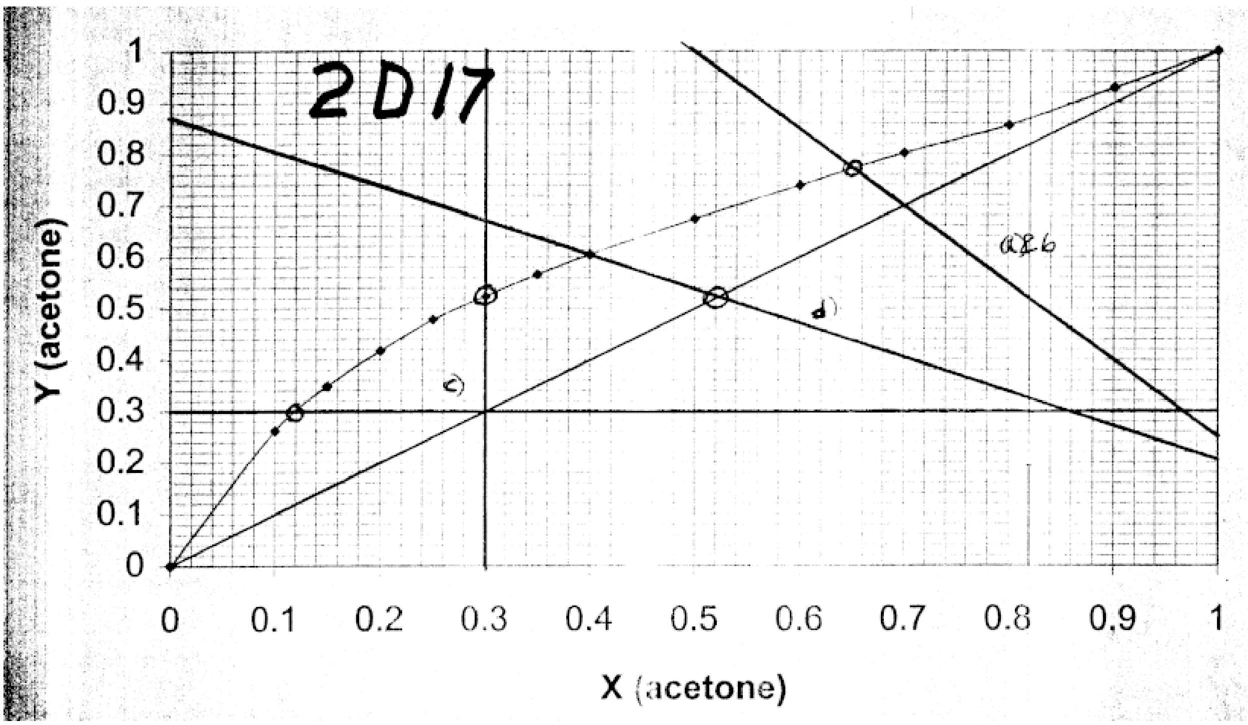
2-D17. a. $V = 0.4F = 400, \quad L = 600 \quad \text{Slope } = -L/F = -1.5$

Intercepts y = x = z = 0.70. Plot line and find $x_A = 0.65$, $y_A = 0.77$ (see graph)

b. V = 2000, L = 3000. Rest identical to part a.

c. Lowest x_A is horizontal op line (L = 0). $x_A = 0.12$

Highest y_A is vertical op line (V = 0). $y_A = 0.52$. See graph



d. V = 600, L = 400, -L/V = -0.667.

Find $x_A = 0.40$ on equilibrium curve. Plot op line & find intersection point with y = x line. $z_A = 0.52$

2.D18. From $x_i = \dfrac{z_i}{1+(K_i-1)\dfrac{V}{F}}$, we obtain $\dfrac{V}{F} = \dfrac{\dfrac{z_h}{x_h}-1}{K_h-1}$

Guess T_{drum}, calculate K_h, K_b and K_p, and then determine V/F.

$$\text{Check: } \sum \frac{(K_1-1)z_i}{1+(K_1-1)V/F} = 0\ ?$$

Initial guess: T_{drum} must be less than temperature to boil pure hexane

$\left(K_h = 1.0,\ T = 94^\circ C\right)$. Try 85°C as first guess (this is not very critical and the calculation will tell us if there is a mistake). $K_h = 0.8,\ K_b = 4.8,\ K_p = 11.7$.

$$\frac{V}{F} = \frac{\frac{0.6}{0.85} - 1}{0.8 - 1} = 1.471 \text{. Not possible. Must have } K_h < \frac{0.6}{0.85} = 0.706$$

Try $T = 73^\circ C$ where $K_h = 0.6$. Then $K_b = 3.8,\ K_p = 9.9$.

$$\frac{V}{F} = \frac{\frac{0.6}{.85} - 1}{.6 - 1} = 0.735$$

Check:

$$\sum \frac{(K_i - 1) z_i}{1 + (K_i - 1) V/F} = \frac{(8.9)(.1)}{1 + (8.9).735} + \frac{(2.8)(.3)}{1 + (2.8).735} + \frac{(-.4)(.6)}{1 - (.4)(735)} = 0.05276$$

Converge on $T \sim 65.6^\circ C$ and $V/F \sim 0.57$.

2.D19. 90% recovery n-hexane means $(0.9)(Fz_{C6}) = L(x_{C6})$

Substitute in $L = F - V$ to obtain $z_{C6}(.9) = (1 - V/F) x_{C6}$

C_8 balance: $z_{C6}F = Lx_{C6} + Vy_{C6} = (F - V) x_{C6} + K_{C6} V x_{C6}$

or $z_{C6} = (1 - V/F) x_{C6} + x_{C6} K_{C6}\, V/F$

Two equations and two unknowns. Remove x_{C6} and solve

$$z_{C6} = .93C_6 + \frac{(.9) z_{C6}\, KV/F}{1 - V/F}$$

Solve for V/F. $\frac{V}{F} = \frac{.1}{(.9K_{C6}) + .1}$. Trial and error scheme.

Pick T, Calc K_{C6}, Calc V/F, and Check $f(V/F) = 0$?

If not $K_{ref_{new}} = \frac{K_{ref}(T_{old})}{1 + d\ f(T)}$

Try $T = 70^\circ C$. $K_{C4} = 3.1,\ K_{C5} = .93,\ K_{C6} = .37 = K_{ref}$

$$\frac{V}{F} = \frac{.1}{(.9)(.37) + .1} = 0.231.$$

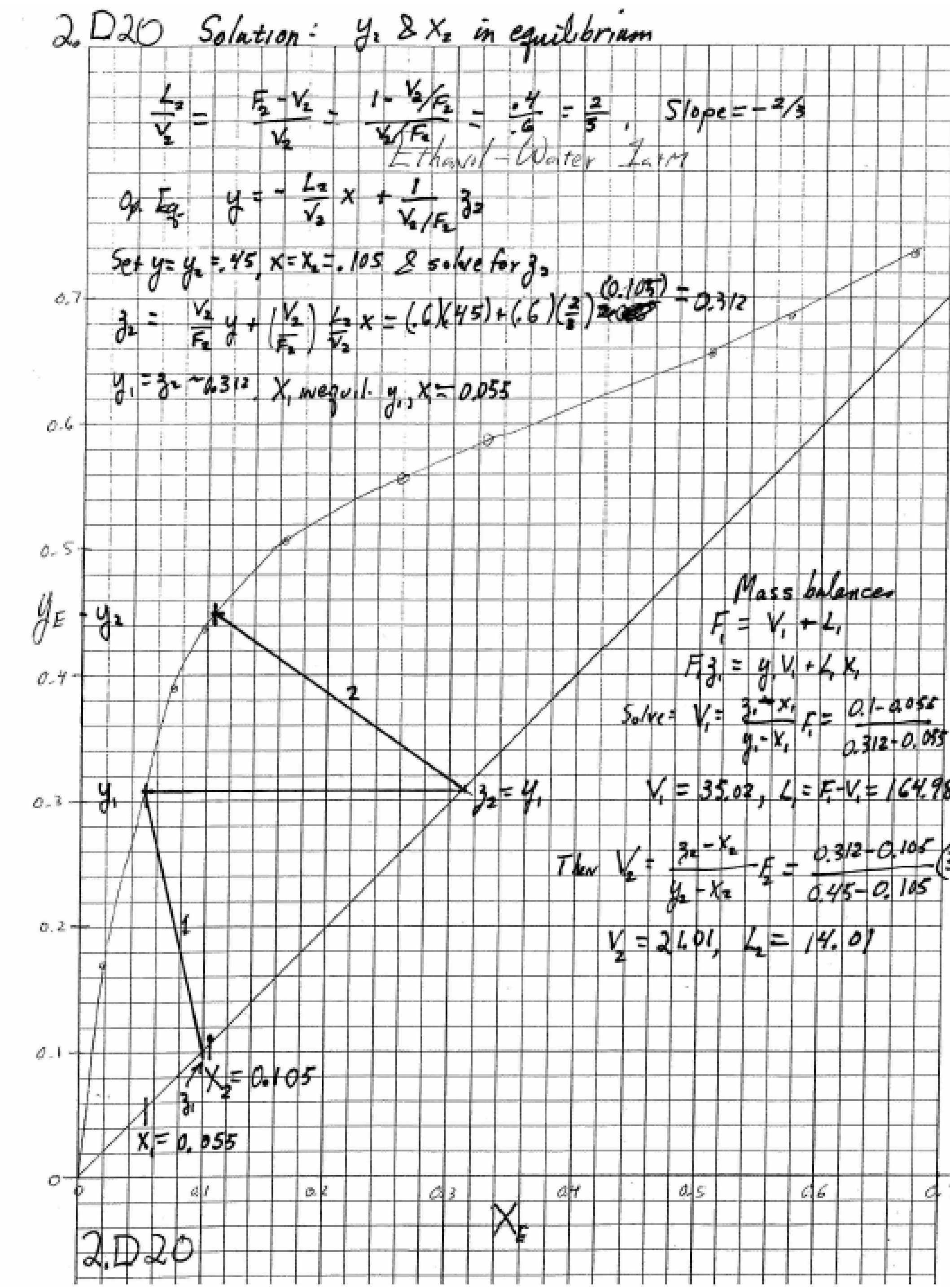
Rachford Rice equation

$$f = \frac{(2.1).4}{1 + (2.1).231} + \frac{(-.08).25}{1 - (.08).231} - \frac{(.63).35}{1 - (.63)(.231)} = .28719$$

$$K_{ref}(T_{new}) = \frac{.37}{1 + 0.28719} = 0.28745\ (\text{use } .28)$$

Converge on $T_{New} \sim 57^\circ C$. Then $K_{C4} = 2.50,\ K_{C8} = .67$, and $V/F = 0.293$.

2.D20. *New Problem 4th ed.*



2.D21. a.) $K_{C2} = 4.8 \quad K_{C5} = 0.153$

Soln to Binary R.R. eq. $\frac{V}{F} = \frac{-z_A}{(K_B - 1)} - \frac{z_B}{(K_A - 1)}$, $\frac{V}{F} = \frac{-0.55}{(.153 - 1)} - \frac{0.45}{(4.8 - 1)} = 0.5309$

$$x_{C2} = \frac{z_{C2}}{1 + (K_{C2} - 1)\frac{V}{F}} = \frac{0.55}{1 + (3.8)(.5309)} = 0.1823, \quad y_{C2} = 0.8749,\ x_{C5} = 0.8177,\quad y_{C5} = 0.1251$$

Need to convert F to kmol. Avg MW $= 0.55(30.07) + 0.45(72.15) = 49.17$

$$F = 100{,}000\,\frac{\text{kg}}{\text{hr}}\left|\frac{\text{kmol}}{49.17\text{ kg}}\right| = 2033.7\text{ kmol/h},\ V = (V/F)F = 1079.7, \quad L = F - V = 954.0\text{ kmol/h}$$

b.) $$u_{Perm} = K_{drum}\sqrt{\frac{\rho_L - \rho_v}{\rho_v}}$$

To find $$\overline{MW}_L = (0.1823)(30.07) + (0.8177)(72.15) = 64.48$$

$$\overline{MW}_V = (0.8749)(30.07) + (0.1251)(72.15) = 35.33$$

For liquid assume ideal mixture:

$$\bar{V}_l = x_{C2}\bar{V}_{C2,liq} + x_{C5}\bar{V}_{C5,liq} = x_{C2}\frac{\overline{MW}_{C2}}{\rho_{C2,liq}} + x_{C5}\frac{\overline{MW}_{C5}}{\rho_{C5,liq}}$$

$$\bar{V}_L = (0.1823)\frac{(30.07)}{0.54} + (0.8177)\frac{(72.15)}{(0.63)} = 103.797\text{ ml/mol}$$

$$\rho_L = \frac{\overline{MW}_L}{\bar{V}_L} = \frac{64.48}{103.797} = 0.621\text{ g/ml}$$

For vapor: ideal gas: $$\rho_v = \frac{\overline{MW}_v}{RT} = \frac{700\text{ kPa}\left|\frac{\text{atm}}{101.3\text{ kPa}}\right|\left(35.33\frac{\text{g}}{\text{mol}}\right)}{82.0575\frac{\text{ml atm}}{\text{mol K}}(303.16\text{K})} = 0.009814\text{ g / ml}$$

$$K_{drum}:\ \text{Use Eq. (2-60) with}\ F_{lV} = \frac{W_L}{W_V}\sqrt{\frac{\rho_v}{\rho_L}}$$

$$W_L = 997.7\frac{\text{kmol}}{\text{h}}\left|\frac{64.48\text{ kg}}{\text{kmol}}\right| = 6{,}4331.7\text{ kg/h},\ W_V = 881.5\left|\underline{35.33}\right| = 31{,}143.4\text{ kg/h}$$

$$F_{lV} = \frac{64331.7}{31{,}143.3}\sqrt{\frac{0.009814}{0.621}} = 0.2597$$

$$K_{drum} = \exp\Big[-1.877478 + (-0.81458)(\ell n\ .2597) + (-0.18707)[\ell n\ 0.2597]^2$$
$$+ (-0.0145229)(\ell n\ 0.2597)^3 + (-0.0010149)(\ell n\ 0.2597)^4\Big] = 0.3372$$

$$u_{Perm} = (0.3372)\sqrt{\frac{0.621 - 0.009814}{0.009814}}\ \frac{\text{ft}}{\text{s}}\left|\frac{1.0\text{ m}}{3.2808\text{ ft}}\right| = 0.8111\text{ m/s}$$

$$A_C = \frac{V\ MW_V}{u_{Perm}\ 3600\ \rho_v} = \frac{\left(1079.7\frac{\text{kmol}}{\text{h}}\right)\left(35.33\frac{\text{kg}}{\text{kmol}}\right)}{0.8111\frac{\text{m}}{\text{s}}\left(3600\frac{\text{s}}{\text{h}}\right)\left(0.009814\frac{\text{g}}{\text{cm}^3}\right)\left(\frac{\text{kg}}{1000\text{g}}\right)\left(\frac{10^6\text{cm}^3}{\text{m}^3}\right)} = 1.392\text{ m}^2$$

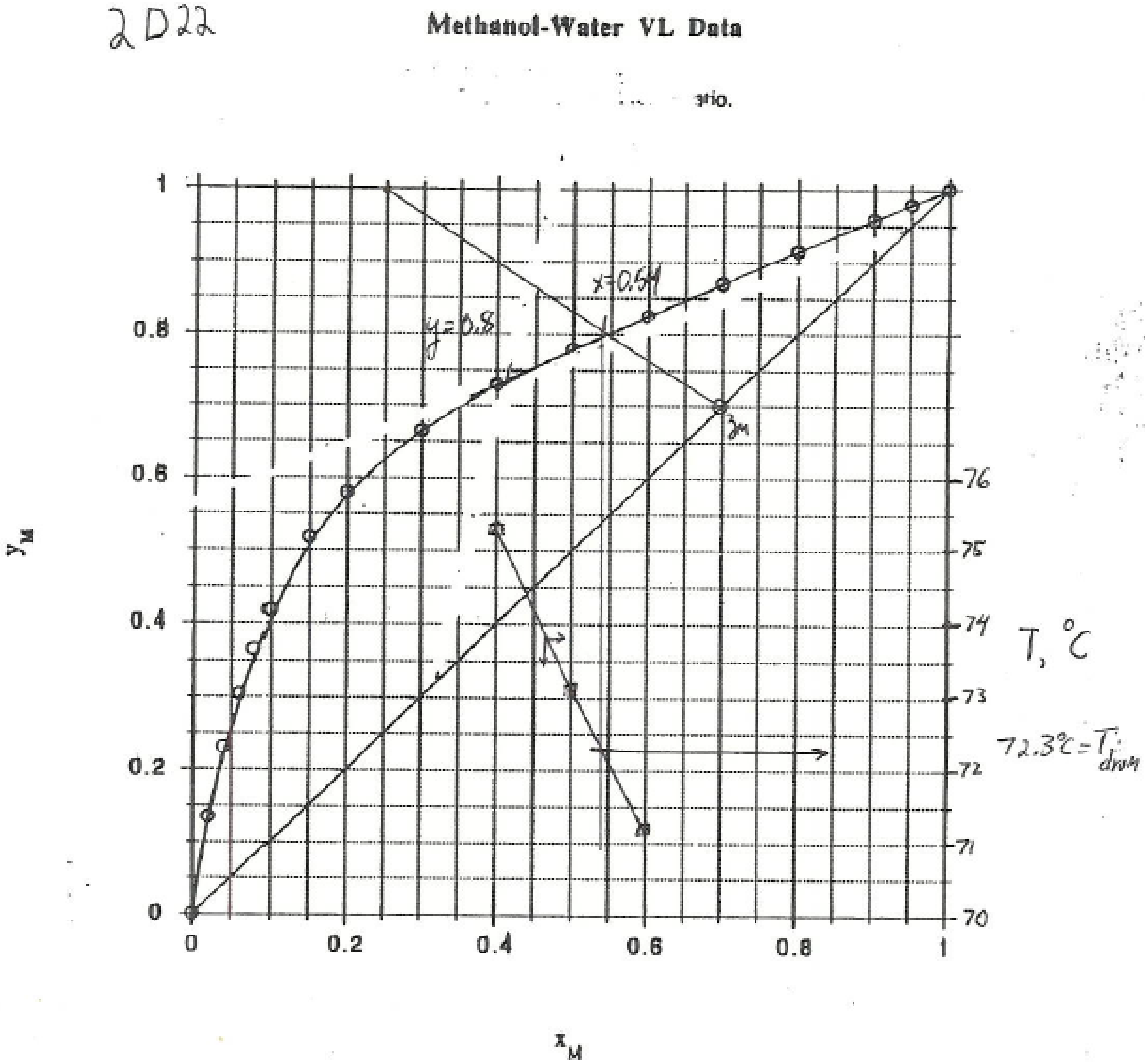
$$D = \sqrt{4A_C/\pi} = 1.33\text{ m}.\ \text{Arbitrarily}\quad L/D = 4, \quad L = 5.32\text{ m}$$

2.D22. *New problem in 4th edition.*

a. V = F – L = 50 – 20 = 20 kmol/h. V/F = 3/5, Slope operating line = -L/V = -20/30 = - 2/3, $z_M = 0.7$

From graph, y = 0.8, x = 0.54.

b. From graph of T vs. x_M, T_{drum} = 72.3°C. (see graph).



2.D23. *New Problem 4th ed.*

Part a. F_{new} = (1500 kmol/h)(1.0 lbmol/(0.45359 kmol)) = 3307 lb mol/h.

V, W_V, L, and W_L are the values in Example 2-4 divided by 0.45359. The conversion factor divides out in F_{lv} term. Thus, F_{lv}, K_{drum}, and u_{perm} are the same as in Example 2-4. The Area increases because V increases: Area = $Area_{Example\ 2\text{-}4}$/0.45359 = 16.047/0.45359 = 35.38 ft^2.

$$Diameter = \sqrt{4Area/\pi} = 6.71\,feet$$

Probably round this off to 7.0 feet and use a drum height of 28 feet.

b. $F_{parallel}$ = 3307 – 1500 = 1807 lbmol/h.

F_{lv}, K_{drum}, and u_{perm} are the same as in Example 2-4. $V_{parallel}$ = (V/F) $F_{parallel}$ = 0.51 (1807) = 921.6 kmol/h.

$$A_c = 16.047 \times \frac{V_{parallel}}{V_{Example\_2-4}} = 16.047 \times (921.6/765) = 19.33\ ft^2$$

Then, $Diameter = \sqrt{4Area/\pi} = 4.96\ feet$, Use a 5.0 feet diameter and a length of 20 feet.

2.D24. p = 300 kPa At any T. $K_{C3} = y_{C3}/x_{C3}$, K's are known. $K_{C6} = y_{C6}/x_{C6} = (1-y_{C3})/(1-x_{C3})$

Substitute 1st equation into 2nd $\quad K_{C6} = (1-K_{C3}x_{C3})/(1-x_{C3})$

Solve for x_{C3}, $\quad (1-x_{C3})K_{C6} = 1-K_{C3}x_{C3}$, $x_{C3}(K_{C3}-K_{C6}) = 1-K_{C6}$

$$x_{C3} = \frac{1-K_{C6}}{K_{C3}-K_{C6}} \quad \& \quad y_{C3} = \frac{K_{C3}(1-K_{C6})}{K_{C3}-K_{C6}}$$

At 300 kPa pure propane $(K_{C3} = 1.0)$ boils at -14°C (Fig. 2-10)

At 300 kPa pure n-hexane $(K_{C6} = 1.0)$ boils at 110°C

Check: at -14°C $\quad x_{C3} = \frac{1-K_{C6}}{1-K_{C6}} = 1, \quad y_{C3} = \frac{1(1-K_{C6})}{1-K_{C6}} = 1.0$

at 110°C $\quad x_{C3} = \frac{0}{K_{C3}} = 0, \quad y_{C3} = \frac{K_{C3}(0)}{K_{C3}} = 0$

Pick intermediate temperatures, find K_{C3} & K_{C6}, calculate x_{C3} & y_{C3}.

T	K_{C3}	K_{C6}	x_{C3}	$y_{C3} = K_{C3}x_{C3}$	
0ºC	1.45	0.027	$\frac{1-0.027}{1.45-0.027} = 0.684$	0.9915	
10ºC	2.1	0.044	0.465	0.976	See
20ºC	2.6	0.069	0.368	0.956	Graph
30ºC	3.3	0.105	0.280	0.924	
40ºC	3.9	0.15	0.227	0.884	
50ºC	4.7	0.21	0.176	0.827	
60ºC	5.5	0.29	0.136	0.75	
70ºC	6.4	0.38	0.103	0.659	

b. $\quad x_{C3} = 0.3$, $V/F = 0.4$, $\quad L/V = 0.6/0.4 = 1.5$

Operating line intersects $\quad y = x = 0.3$, $\quad$ Slope -1.5

$$y = -\frac{L}{V}x + \frac{F}{V}z \quad \text{at} \quad x = 0, \quad y = \frac{F}{V}z = \frac{0.3}{0.4} = 0.75$$

Find $y_{C3} = 0.63$ and $x_{C3} = 0.062$

Check with operating line: $0.63 = -1.5(.062) + 0.75 = 0.657$ OK within accuracy of the graph.

c. Drum T: $K_{C3} = y_{C3}/x_{C3} = 0.63/0.062 \approx 10.2$, DePriester Chart T = 109ºC

d. $y = .8, \quad x \sim .16 \quad$ $\text{Slope} = -\frac{L}{V} = \frac{\Delta y}{\Delta x} = \frac{.8-.6}{.16-.6} = -0.45 \quad = -\frac{1-f}{f} = -.45$

V/F = f =1/1.45 = 0.69

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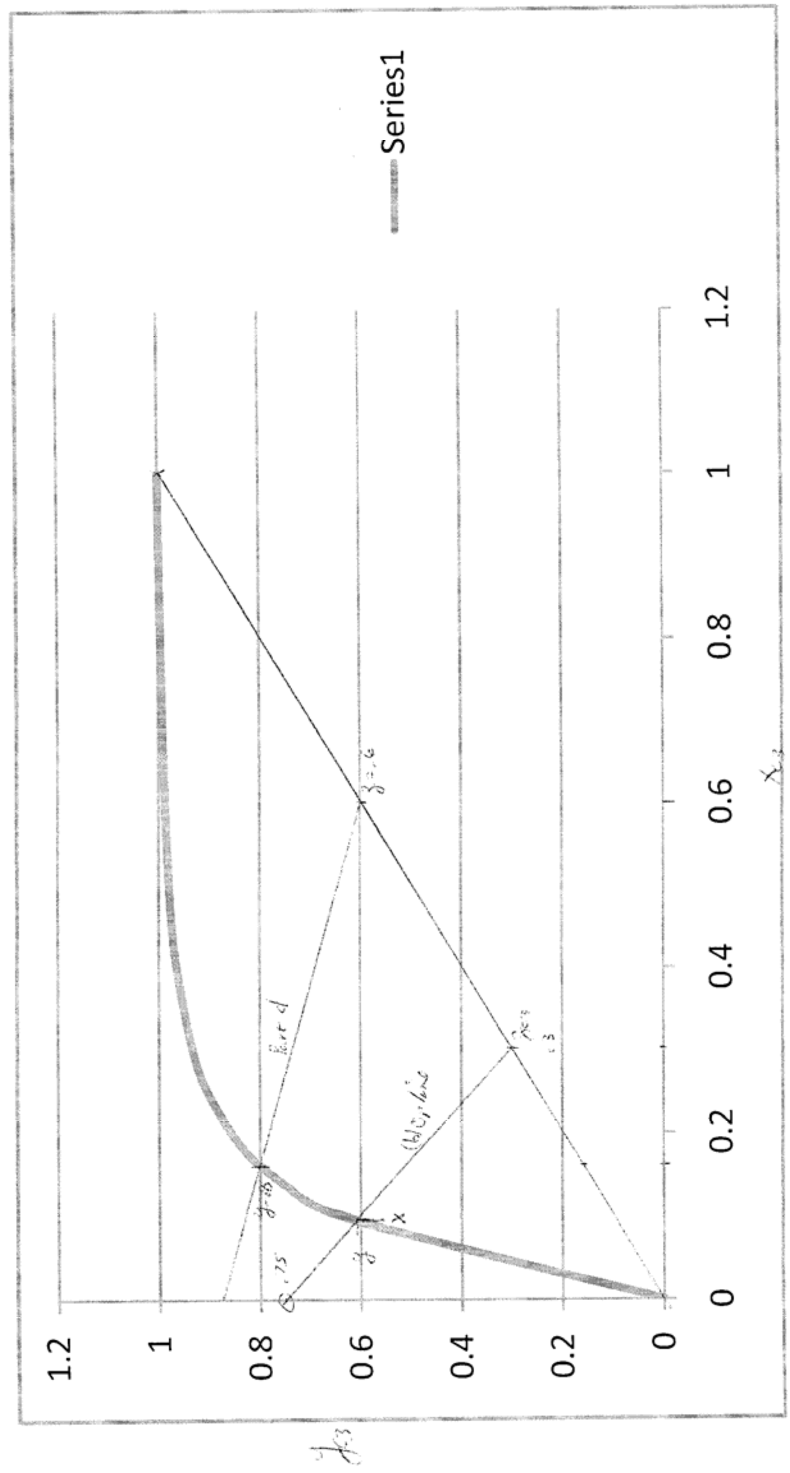
Prob 1



2.D25. 20% Methane and 80% n-butane. $T_{drum} = .50\,^{o}C$, $\frac{V}{F} = 0.40$, Find p_{drum}

$$0 = f\left(\frac{V}{F}\right) = \frac{(K_A - 1)z_A}{1 + (K_A - 1)\left(\frac{V}{F}\right)} + \frac{(K_B - 1)z_B}{1 + (K_B - 1)\frac{V}{F}}$$

Pick $p_{drum} = 1500$ kPa: $K_{C4} = 13$ $\quad K_{nC4} = 0.4$

(Any pressure with $K_{C1} > 1$ and $K_{C4} < 1.0$ is OK)

Trial 1 $\quad f_1 = \frac{12(.2)}{1 + 12(.4)} + \frac{(-.6)(.8)}{1 - .6(.4)} = -0.2178$ $\quad$ Need lower p_{drum}

$$K_{C4}(P_{new}) = \frac{K_{C4}(P_{old})}{1 + (d) f(P_{old})} = \frac{(0.4)}{1 + (-.2138)} = 0.511 \text{ with d = 1.0}$$

$$P_{new} = 1160 \quad K_{C1} = 16.5,\ f_2 = \frac{(15.5)(.2)}{1 + 15.5(.4)} + \frac{-.489(.8)}{1 - (.489)(.4)} = 0.4305 + -.4863 = -0.055769$$

$$K_{C4}(P_{new}) = \frac{0.511}{1 + -0.055769} = 0.541,\ \ P_{new} = 1100,\ K_{C1} = 17.4$$

$$f_3 = \frac{16.4(.2)}{1 + (16.4)(.4)} + \frac{(-.459)(.8)}{1 - (.459)(.4)} = -0.0159, \text{ OK. Drum pressure = 1100 kPa}$$

b.) $\quad x_i = \frac{z_i}{1 + (K_i - 1)\frac{V}{F}},\ \ x_{C1} = \frac{0.2}{1 + (16.4)(.4)} = 0.02645$

$$y_{C1} = K_{C1} x_{C1} = (17.4)(0.02645) = 0.4603$$

2.D26. a) Can solve for L and V from M.B. $\quad 100 = F = V + L$

$45 = Fz = 0.8V + 0.2162L$

Find: $\quad$ L = 59.95 and V = 40.05

b) Stage is equil. $\quad K_{C3} = \frac{y_{C3}}{x_{C3}} = \frac{0.8}{0.2162} = 3.700,\ \ K_{C5} = \frac{0.2}{0.7838} = .2552$

These K values are at same T, P. Find these 2 K values on DePriester chart. Draw straight line between them. Extend to T_{drum}, p_{drum}. Find 10ºC, 160 kPa.

2.D27. a.) VP_{C5}: $\log_{10} VP = 6.853 - \frac{1064.8}{0 + 233.01} = 2.2832,\ VP = 191.97$ mmHg

b.) $VP = 3 \times 760 = 2280$ mmHg, $\log_{10} VP = (6.853) - 1064.8/(T + 233.01)$

Solve for $\quad$ T = 71.65ºC

c.) $\quad P_{tot} = 191.97$ mm Hg [at boiling for pure component $P_{tot} = VP$]

d.) C5: $\log_{10} VP = 6.853 - \frac{1064.8}{30 + 233.01} = 2.8045,\ \ VP = 637.51$ mm Hg

$$K_{C5} = VP_{C5}/P_{tot} = 637.51/500 = 1.2750$$

C6: $\log_{10} VP_{C6} = 6.876 - \dfrac{1171.17}{30 + 224.41} = 2.2725$, $VP_{C6} = 187.29$ mm Hg

$$K_{C6} = 187.29/500 = 0.3746$$

e.) $K_A = y_A/x_A$ $\quad K_B = y_B/x_B = (1-y_A)/(1-x_A)$

If K_A & K_B are known, two eqns. with 2 unknowns $(K_A \ \& \ y_A)$ Solve.

$$x_{C5} = \frac{1-K_{C6}}{K_{C5}-K_{C6}} = \frac{1-0.3746}{1.2750-0.3746} = 0.6946$$

$$y_{C5} = K_{C5}x_{C5} = (1.2750)(0.6946) = 0.8856$$

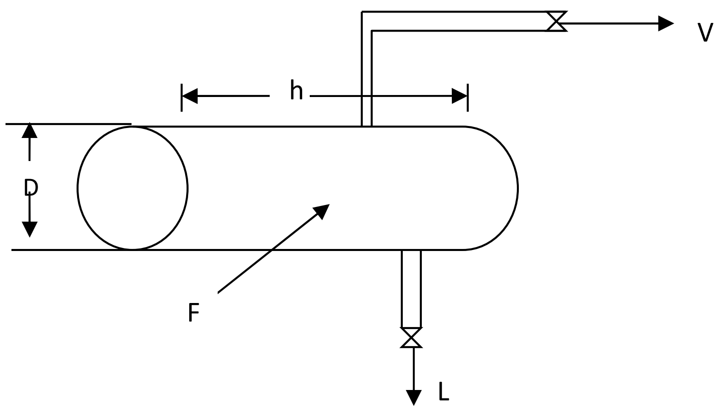
f.) Overall, M.B., F = L + V or 1 = L + V

$C5: \ Fx_F = Lx + Vy \quad .75 = 0.6946\ L + 0.8856\ V$

Solve for L & V: L = 0.7099 & V = 0.2901 mol

g.) Same as part f, except units are mol/min.

2.D28.



From example 2-4, $x_H = 0.19$, $T_{drum} = 378K$, $V/F = 0.51$, $y_H = 0.6$, $z_H = 0.40$

MW_v = 97.39 lbm/lbmole (Example 2-4)

$$\rho_v = 3.14\times10^{-3}\ g/mol \left| \frac{1}{454 g/lbm} \right| \frac{28316.85 cm^3}{ft^3} = 0.198 \frac{lbm}{ft^3}$$

Example 2.4

$$u_{perm} = K_{drum}\sqrt{\frac{\rho_L - \rho_v}{\rho_V}}, \qquad K_{horiz} = 1.25\ K_{vertical}$$

From Example 2-4, $K_{vertical} = 0.4433$, $K_{horiz} = 1.25(0.4433) = 0.5541$

$$u_{perm} = 0.5541\left(\frac{0.6960-0.00314}{0.00314}\right)^{1/2} = 8.231\ ft/s$$ [densities from Example 2-4]

$$V = \left(\frac{V}{F}\right)F = (0.51)\left(3000\frac{\text{lbmol}}{\text{h}}\right) = 1530 \text{ lbmol/h}$$

$$A_{vap} = \frac{1530\frac{\text{lbmol}}{\text{h}}\left(97.39\frac{\text{lbm}}{\text{lbmole}}\right)}{\left(8.231\frac{\text{ft}}{\text{s}}\right)\left(3600\frac{\text{s}}{\text{h}}\right)\left(0.1958\frac{\text{lbm}}{\text{ft}^3}\right)} = 25.68 \text{ ft}^2$$

$$A_{total} = A_{vap}/0.2 = 128.4\text{ft}^3, D_{min} = \sqrt{4A_{total}/\pi} = 12.8\text{ft}$$

$$V_{liq} = \frac{160,068}{43.41}\frac{55+85}{60\,\text{min/h}} = 8603.8\text{ft}^3, h = \frac{5V_{liq}}{\pi D^2} = 83.51\text{ft} \text{ and h/D = 6.5.}$$

2.D29. The stream tables in Aspen Plus include a line stating the fraction vapor in a given stream. Change the feed pressure until the feed stream is all liquid (fraction vapor = 0). For the Peng-Robinson correlation the appropriate pressure is 74 atm.
The feed *mole* fractions are: methane = 0.4569, propane = 0.3087, n-butane = 0.1441, i-butane = 0.0661, and n-pentane = 0.0242.
b. At 74 atm, the Aspen Plus results are; L = 10169.84 kg/h = 201.636 kmol/h, V = 4830.16 kg/h = 228.098 kmol/h, and T_{drum} = -40.22 ºC.
The vapor mole fractions are: methane = 0.8296, propane = 0.1458, n-butane = 0.0143, i-butane = 0.0097, and n-pentane = 0.0006.
The liquid mole fractions are: methane = 0.0353, propane = 0.4930, n-butane = 0.2910, i-butane = 0.1298, and n-pentane = 0.0509.

c. Aspen Plus gives the liquid density = 0.60786 g/cc, liquid avg MW = 50.4367, vapor density = 0.004578 g/cc = 4.578 kg/m^3, and vapor avg MW = 21.17579 g/mol = kg/kmol.
The value of u_{perm} (in ft/s) can be determined by combining Eqs. (2-64), (2-65) and (2-69)

$F_{lv} = (W_L/W_V)[\rho_V/\rho_L]^{0.5} = (10169.84/4830.16)[0.004578/0.60786]^{0.5} = 0.18272$

Resulting $K_{vertical}$ = 0.378887 , $K_{horizontal}$ = 0.473608, and u_{perm} = 5.436779 ft/s = 1.657m/s

$$A_{vap} = \frac{4830.16\frac{\text{kg}}{\text{h}}}{\left(1.657\frac{\text{m}}{\text{s}}\right)\left(3600\frac{\text{s}}{\text{h}}\right)\left(4.578\frac{\text{kg}}{\text{m}^3}\right)} = 0.177 \text{ m}^2$$

$$A_{total} = A_{vap}/0.2 = 0.884\text{m}^3, D_{min} = \sqrt{4A_{total}/\pi} = 1.06\text{m}$$

$$h/D = 6 = \frac{5V_{liq}}{\pi D^2}, \text{ thus } V_{liq} = 6\pi D^2/5 = 4.23\text{m}^3$$

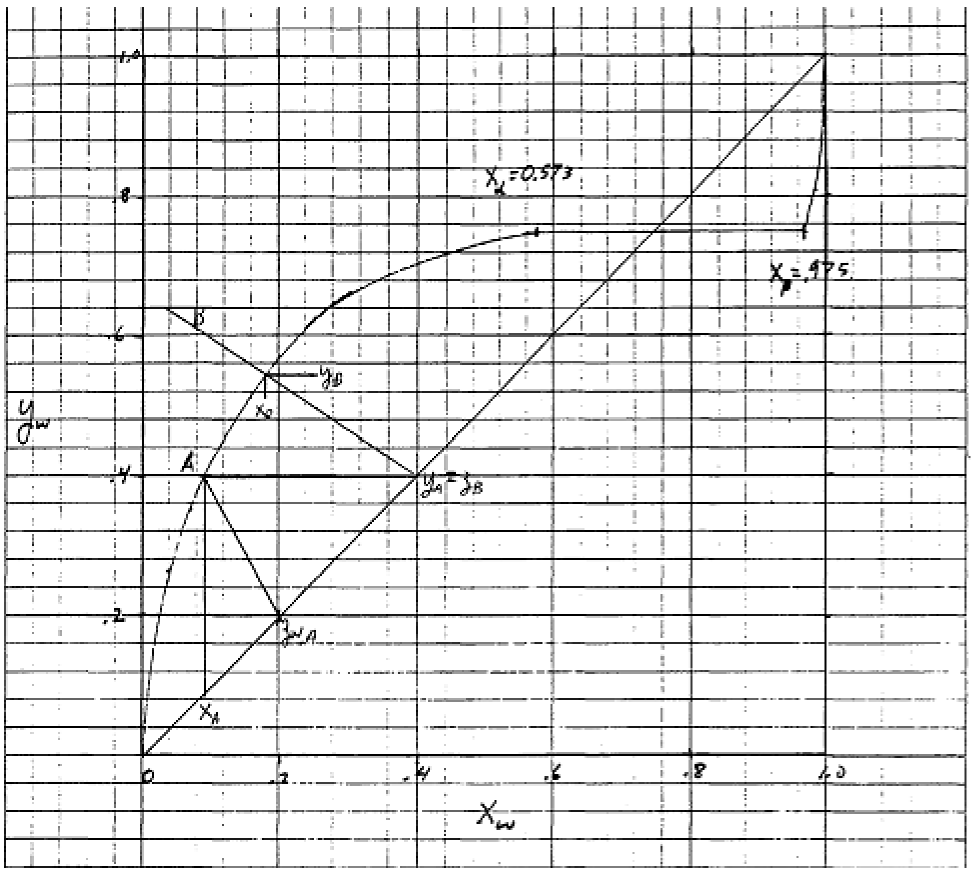
$$V_{liq} = (\text{Vol rate})(\text{hold time} + \text{surge time}) = \left(\frac{10169.84\text{kg/h}}{607.86\text{kg/m}^3}\right)(9/60 + st)$$

$$st = 607.86V_{liq}/10169.84 - 9/60 = 0.103 \text{ hours} = 6.18 \text{ min}$$

2.D30.. a. From the equilibrium data if y_A = .40 mole fraction water, then x_A = 0.09 mole fraction water. Can find L_A and V_A by solving the two mass balances for stage A simultaneously.
$L_A + V_A = F_A = 100$ and $L_A(.09) + V_A(.40) = (100)(.20)$. The results are V_A = 35.48 and L_A = 64.52.

b. In chamber B, since 40 % of the vapor is condensed, $(V/F)_B = 0.6$. The operating line for this flash chamber is,
$y = -(L/V)x + F_B/V)\ z_B$ where $z_B = y_A = 0.4$ and $L/V + .4F_B/.6F_B = 2/3$. This operating line goes through the point $y = x = z_B = 0.4$ with a slope of -2/3. This is shown on the graph. Obtain $x_B = 0.18$ & $y_B = 0.54$. L_B = (fraction condensed)(feed to B) = 0.4(35.48) = 14.19 kmol/h and $V_B = F_B - L_B = 21.29$.

c. From the equilibrium if $x_B = 0.20$, $y_B = 0.57$. Then solving the mass balances in the same way as for part a with $F_B = 35.48$ and $z_B = 0.4$, $L_B = 16.30$ and $V_B = 19.18$. Because $x_B = z_A$, recycling L_B does not change $y_B = 0.57$ or $x_A = 0.09$, but it changes the flow rates $V_{B,new}$ and $L_{A,new}$. With recycle these can be found from the overall mass balances: $F = V_{B,new} + L_{A,new}$ and $Fz_A = V_{B,new}y_B + L_{A,new}\ x_A$. Then $V_{B,new}$ = 22.92 and $L_{A,new}$ = 77.08.



Graph for problem 2.D30.

2.D31. *New problem in 4th US edition. Was 2.D13 in 3rd International Edition.*

a) Since K's are for mole fractions, need to convert feed to mole fractions.

Basis: 100 kg feed

$$50\ \text{kg n C}_4 \frac{1\ \text{kmol}}{58.12\ \text{kg}} = 0.8603\ \text{kmol} \qquad z_4 = 0.555$$

$$50\ \text{kg n C}_5 \frac{1\ \text{kmol}}{72.5\ \text{kg}} = 0.6897\ \text{kmol n C}_5 \qquad z_5 = 0.445$$

Total 1.5499 kmol

DePriester Chart $\quad K_{C4} = 2.05, \ K_{C5} = 0.58, \ \text{(Result similar if use Raoult's law)}.$

$$\frac{V}{F} = \frac{-0.555}{0.58-1} - \frac{0.445}{1.05} = 1.3214 - 0.424 = 0.8976$$

Check $\quad f\left(\frac{V}{F}\right) = \frac{(1.05)(.555)}{1+(.05)(.8976)} + \frac{(-.42).445}{1-.42(.8976)} = 0.3000 - .29999 = 0 \ \text{OK}$

Eq. 3.23 $\quad x_{C4} = \frac{z_{C4}}{1+(K_{C4}-1)V/F} = \frac{.555}{1+1.056(.8976)} = 0.2857$,

$$x_{C5} = .7143 \quad y_{C4} = K_{C4}x_{C4} = 0.5857 \ , \ y_{C5} = 0.4143$$

b) From problem 2.D.g., $K_{C4} = 1.019$ and $K_{C5} = 0.253$.
Solving RR equation,

$$\frac{V}{F} = \left[-\frac{z_A}{(K_B-1)} - \frac{z_B}{(K_A-1)} \right] = \frac{-.555}{(0.253-1)} - \frac{0.445}{0.019} = -23.28$$

NOT possible. Won't flash at 0ºC.

2.D32. *New problem in 4th US edition. Was 2.D28 in 3rd International Edition.*

$(V/F)_A = 2/3, \quad L/V = \frac{1/3}{2/3} = 1/2 \quad$ Slope $= -1/2$ Through $y = x = z_A = 0.6$ See figure

a. $L_A = \frac{1}{3}F = 33.33, \ x_{M,A} = 0.375$ (from Figure) $V_A = \frac{2}{3}F = 66.67, \quad y_{M,A} = 0.72$ (from Figure)

$$\left.\frac{V}{F}\right|_B = 0.4 \quad -\left(\frac{L}{V}\right)_B = -\frac{1-f}{f} = -\frac{0.6}{0.4} = -1.5$$

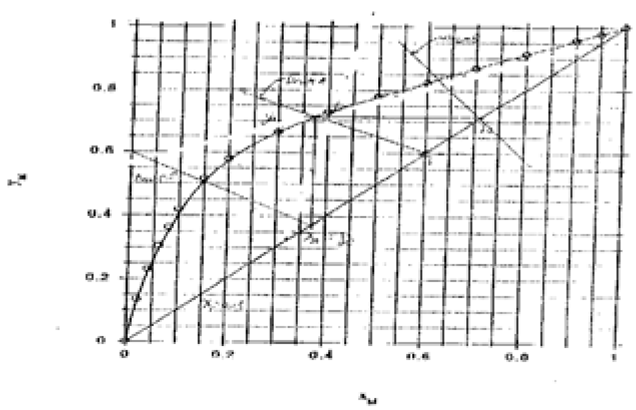
Through $y = x = z_B = y_A = 0.72$

$$V_B = 0.4F_B = 0.4V_A = 0.4(66.67) = 26.67, \quad L_B = 0.6F_B = 0.6(66.67) = 40.00$$

b. $z_C = x_A = 0.375, \quad x_C = 0.15, \quad F_C = L_A = 33.33$, From equilibrium $y_C = 0.51$

At $x = 0, \ y_C = 0.60 = \left(\frac{F}{V}\right)_C z_C \Rightarrow \left(\frac{V}{F}\right)_C = \frac{z_C}{y_C} = \frac{0.375}{0.6} = 0.625$

$$V_C = \left(\frac{V}{F}\right)_C F_C = 0.625(33.33) = 20.83, \quad L_C = F_C - V_C = 33.33 - 20.83 = 12.5$$



2.E1. From Aspen Plus run with 1000 kmol/h at 1 bar, L = V = 500 kmol/h, W_L = 9212.78 kg/h, W_V = 13010.57 kg/h, liquid density = 916.14 kg/m^3 , liquid avg MW = 18.43, vapor density = 0.85 kg/m^3 , and vapor avg MW = 26.02, T_{drum} = 94.1 ^oC, and Q = 6240.85 kW.

The diameter of the vertical drum in meters (with u_{perm} in ft/s) is

$D = \{[4(MW_V)\ V]/[3600\ \pi\ \rho_V\ u_{perm}\ (1\ m/3.281\ ft)]\}^{0.5} =$
$\{[4(26.02)(500)]/[3600(3.14159)(0.85)(1/3.281)u_{perm}]\}^{0.5}$

$F_{lv} = (W_L/W_V)[\rho_V/\ \rho_L]^{0.5} = (9212.78/13010.57)[0.85/916.14]^{0.5} = 0.02157$

Resulting $K_{vertical}$ = 0.404299, and u_{perm} = 13.2699 ft/s, and D = 1.16 m. Appropriate standard size would be used. Mole fractions isopropanol: liquid = 0.00975, vapor = 0.1903

b. Ran with feed at 9 bar and p_{drum} at 8.9 bar with V/F = 0.5. Obtain W_L = 9155.07 kg/h, W_V = 13068.27, density liquid = 836.89, density vapor = 6.37 kg/m^3

$D = \{[4(MW_V)\ V]/[3600\ \pi\ \rho_V\ u_{perm}\ (1\ m/3.281\ ft)]\}^{0.5} =$
$\{[4(26.14)(500)]/[3600(3.14159)(6.37)(1/3.281)u_{perm}]\}^{0.5}$

$F_{lv} = (W_L/W_V)[\rho_V/\ \rho_L]^{0.5} = (9155.07/13068.27)[6.37/836.89]^{0.5} = 0.06112$

Resulting $K_{vertical}$ = .446199, u_{perm} = 5.094885 ft/s, and D = 0.684 m. Thus, the method is feasible.

c. Finding a pressure to match the diameter of the existing drum is trial and error. If we do a linear interpolation between the two simulations to find a pressure that will give us D = 1.0 m (if linear), we find p = 3.66. Running this simulation we obtain, W_L = 9173.91 kg/h, W_V = 13049.43, density liquid = 874.58, density vapor = 2.83 kg/m^3, MW_v = 26.10

$D = \{[4(MW_V)\ V]/[3600\ \pi\ \rho_V\ u_{perm}\ (1\ m/3.281\ ft)]\}^{0.5} =$
$\{[4(26.10)(500)]/[3600(3.14159)(2.83)(1/3.281)u_{perm}]\}^{0.5}$

$F_{lv} = (W_L/W_V)[\rho_V/\ \rho_L]^{0.5} = (9173.91/13049.43)[2.83/874.58]^{0.5} = 0.0400$

Resulting $K_{vertical}$ = .441162, u_{perm} = 7.742851 ft/s, and D = 0.831 m.

Plotting the curve of D versus p_{drum} and setting D = 1.0, we interpolate p_{drum} = 2.1 bar At p_{drum} = 2.1 bar simulation gives, W_L = 9188.82 kg/h, W_V = 13034.53, density liquid = 893.99 , density vapor = 1.69 kg/m^3, MW_v = 26.07.

$D = \{[4(MW_V)\ V]/[3600\ \pi\ \rho_V\ u_{perm}\ (1\ m/3.281\ ft)]\}^{0.5} =$
$\{[4(26.07)(500)]/[3600(3.14159)(1.69)(1/3.281)u_{perm}]\}^{0.5}$

$F_{lv} = (W_L/W_V)[\rho_V/\ \rho_L]^{0.5} = (9188.82/13034.53)[1.69/893.99]^{0.5} = 0.0307$

Resulting $K_{vertical}$ = .42933, u_{perm} = 9.865175ft/s, and D = 0.953 m.

This is reasonably close and will work OK. T_{drum} = 115.42 ^oC, Q = 6630.39 kW,
Mole fractions isopropanol: liquid = 0.00861, vapor = 0.1914

In this case there is an advantage operating at a somewhat elevated pressure.

2.E2. This problem was 2.D13 in the 2nd edition of *SPE*.

a. Will show graphical solution as a binary flash distillation. Can also use R-R equation. To generate equil. data can use

$$x_{C6} + x_{C8} = 1.0, \text{ and } y_{C6} + y_{C8} = K_{C6}x_{C6} + K_{C8}x_{C8} = 1.0$$

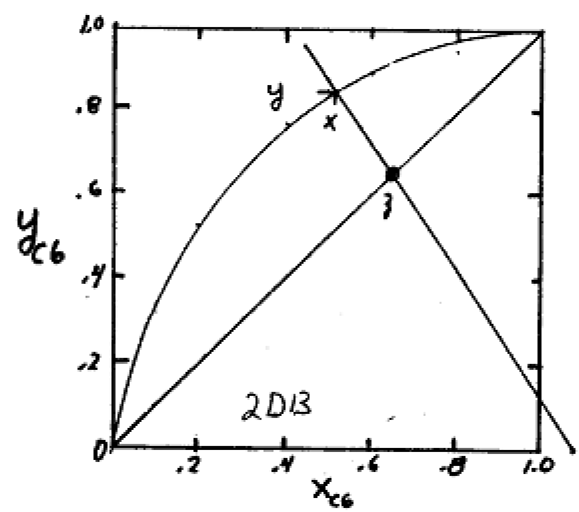
Substitute for x_{C6} $\quad x_{C6} = \dfrac{1 - K_{C8}}{K_{C6} - K_{C8}}$

Pick T, find K_{C6} and K_{C8} (e.g. from DePriester charts), solve for x_{C6}. Then $y_{C6} = K_{C6}x_{C6}$

T°C	K_{C6}	K_{C8}	x_{C6}	$y_{C6} = K_{C6}\, x_{C6}$
125	4	1.0	0	0
120	3.7	.90	.0357	.321
110	3.0	.68	.1379	.141
100	2.37	.52	.2595	.615
90	1.8	.37	.4406	.793
80	1.4	.26	.650	.909
66.5	1.0	.17	1.0	1.0

Op Line Slope $= -\dfrac{L}{V} = -\dfrac{1 - V/F}{V/F} = -\dfrac{.6}{.4} = -1.5$, Intersection y = x = z = 0.65.

See Figure. $y_{C6} = 0.85$ and $x_{C6} = 0.52$. Thus $K_{C6} = .85/.52 = 1.63$.
This corresponds to T = 86°C = 359K



b. Follows Example 2-4.

$$\overline{MW}_L = x_{C8}\left(MW\right)_{C6} + x_{C8}\left(MW\right)_{C8} = (.52)(86.17) + (.48)(114.22) = 99.63$$

$$\bar{V}_L = x_{C6}\frac{(MW)_{C6}}{\rho_{C6}} + x_{C8}\frac{(MW)_{C8}}{\rho_{C8}} = (.52)\frac{86.17}{.659} + (.48)\frac{114.22}{.703} = 145.98 \text{ ml/mol}$$

$$\rho_L = \frac{\overline{MW}_L}{\bar{V}_L} = \frac{99.63}{145.98} = .682 \text{ g/ml} \left[\frac{28316 \text{ ml/ft}^3}{454 \text{ g/lbm}}\right] = 42.57\frac{\text{lbm}}{\text{ft}^3}$$

$$\overline{MW}_v = y_{C6}(MW_{C6}) + y_{C8}(MW_{C8}) = .85(86.17) + .15(114.22) = 90.38$$

$$\rho_v = \frac{p\overline{MW}_v}{RT} = \frac{(1.0)90.38 \text{ g/mol}}{82.0575 \ \frac{\text{ml atm}}{\text{mol}\cdot\text{K}}(359\text{K})} = 0.00307 \text{ g/ml } = 0.19135 \text{ lbm/ft}^3$$

Now we can determine flow rates

$$V = \left(\frac{V}{F}\right)F = (.4)(10,000) = 4000 \text{ lbmol/h}$$

$$W_v = V\left(\overline{MW}_v\right) = 4000(90.38) = 361,520 \text{ lb/h}$$

$$L = F - V = 6000 \text{ lbmol/h}, \ W_L = L\left(\overline{MW}_L\right) = (6000)(99.63) = 597,780 \text{ lb/h}$$

$$F_{lv} = \frac{W_L}{W_v}\sqrt{\frac{\rho_v}{\rho_L}} = \frac{597,780}{361,520}\sqrt{\frac{0.19135}{42.57}} = 0.111, \ \ell n F_{lv} = -2.1995$$

$$K_{drum} = \exp\Big[(-1.87748) + (-.81458)(-2.1995) + (-.18707)(-2.1995)^2$$

$$+ \ (-0.01452)(-2.1995)^3 + (-0.00101)(-2.1995)^4\Big] = 0.423$$

$$u_{Perm} = K_{drum}\sqrt{\rho_L - \rho_v/\rho_v} = (0.423)\sqrt{(42.57 - 19135)/.19135} = 6.30 \text{ ft/s}$$

$$A_{Cs} = \frac{V\left(\overline{MW}_v\right)}{u_{Perm}(3600)\rho_v} = \frac{(4000)(90.38)}{(6.3)(3600)(0.19135)} = 83.33 \text{ ft}^2$$

$$D = \sqrt{4A_{Cs}/\pi} = \sqrt{4(83.33)/\pi} = 10.3 \text{ ft. Use 10.5 ft.}$$

L ranges from 3 × 10.5 = 31.5 ft to 5 × 10.5 = 52.5 ft.

Note: This u_{Perm} is at 85% of flood. If we want to operate at lower % flood (say 75%)

$$u_{Perm_{75\%}} = (0.75/0.85)u_{Perm_{85\%}} = (0.75/0.85)(.63) = 5.56$$

Then at 75% of flood, A_{Cs} = 94.44 which is D = 10.96 or 11.0 ft.

2.E3. *New problem 4th edition.* The difficulty of this problem is it is stated in weight units, but the VLE data is in molar units. The easiest solution path is to work in weight units, which requires converting some of the equilibrium data to weight units and replotting – good practice. The difficulty with trying to work in molar units is the ratio L/V = 0.35/0.65 = 0.5385 in weight units becomes in molar units,

$\frac{L_{molar}}{V_{molar}} = \frac{L_{wt}}{V_{wt}}\frac{(MW)_{vapor}}{(MW)_{liquid}}$, but x and y are not known the molecular weights are unknown.

In weight units, V = F(V/F) =2000 kg/h (0.35) = 700 kg/h. L = F – V = 1300 kg/h.

In weight units the equilibrium data (Table 2-7) can be converted as follows:

Basis: 1 mol, x = 0.4 and y = 0.729, T = 75.3 C

Liquid: 0.4 mol methanol ×32.04 g/mol = 12.816 g

0.6 mol water × 18.016 g/mol = 10.806 g

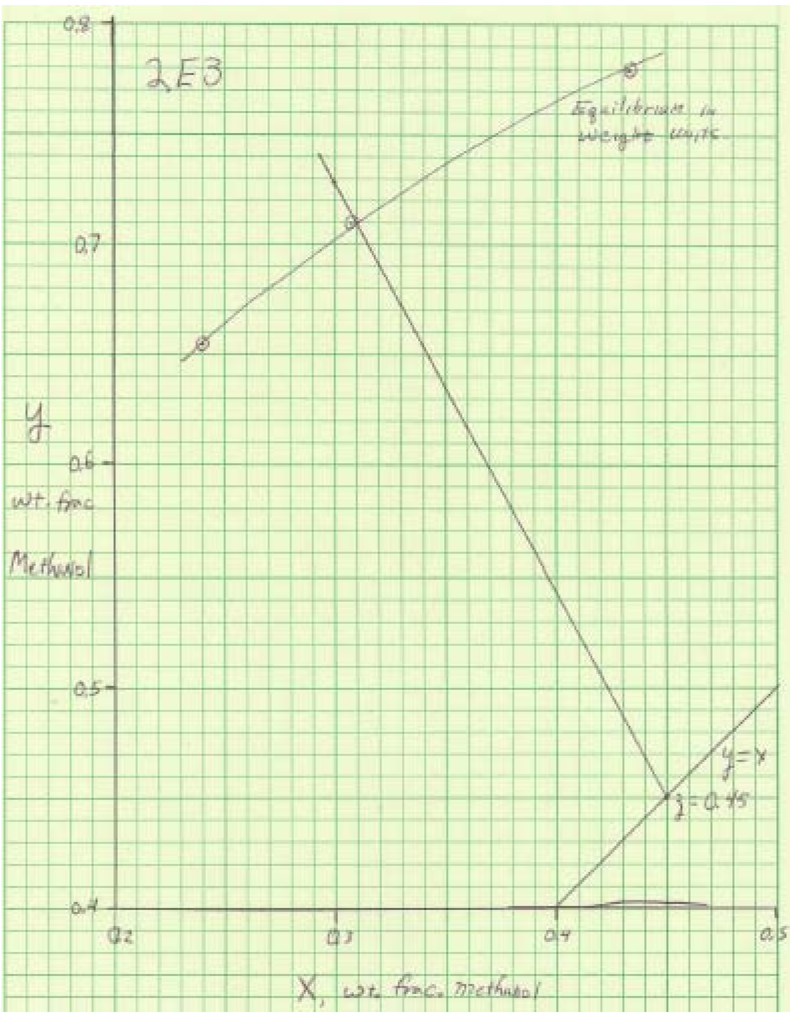
Total = 23.622 g → x = 0.5425 wt frac methanol

Vapor: 0.729 mol methanol = 23.357 g

0.271 mol water = 4.881 g

28.238 g → y = 0.8271 wt frac methanol.

Similar calculations for: 0.3 mole frac liquid give $x_{wt} = 0.433$ and $y_{wt} = 0.7793$, T = 78.0 C
0.2 mole frac liquid give $x_{wt} = 0.3078$ and $y_{wt} = 0.7099$, T = 81.7 C
0.15 mole frac liquid give $x_{wt} = 0.2389$ and $y_{wt} = 0.6557$, T = 84.4 C.
Plot this data on y_{wt} vs x_{wt} diagram. Operating line is $y = -(L/V)x + (F/V)z$ in weight units.
Slope = - 1.857, y =x = z = 0.45, and y intercept = z/(V/F) = 1.286 all in weight units.
Result is $x_{M,wt} = 0.309$, $y_{M,wt} = 0.709$ (see graph). Note that plotting only the part of the graph needed to solve the problem, the scale could be increased resulting in better accuracy. By linear interpolation $T_{drum} = 81.66$ C.

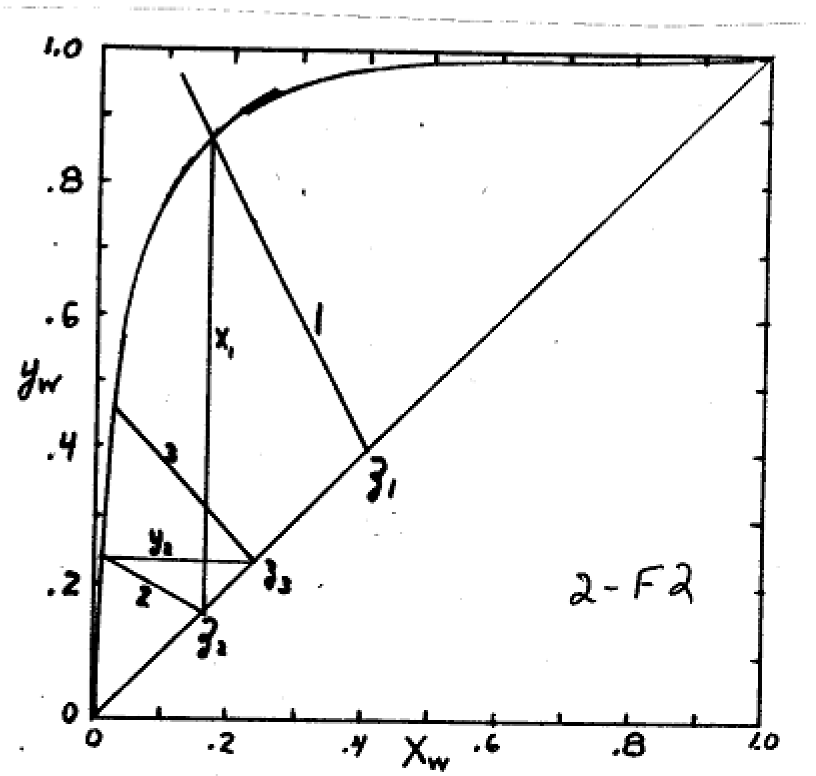


2.F1

x_B	0	.1	.2	.3	.4	.5	.6	.7	.8	.9	1
y_B	0	.22	.38	.52	.62	.71	.79	.85	.91	.96	1

Benzene-toluene equilibrium is plotted in Figure 13-8 of *Perry's Chemical Engineers Handbook*, 6th ed.

2.F2. See Graph. Data is from *Perry's Chemical Engineers Handbook*, 6th ed., p. 13-12.



Stage 1) $\quad z_{F_1} = .4 \quad f = 1/3 \quad \text{Slope} = -\frac{2/3}{1/3} = -2,$

$$\text{Intercept} = \frac{.4}{1/3} = 1.2 \quad y_1 = .872 \quad x_1 = .164 = z_2$$

Stage 2) $\quad z_{F_2} = .164 \quad f = 2/3 \quad \text{Slope} = -\frac{1/3}{2/3} = -1/2$

$$\text{Intercept} = \frac{.164}{2/3} = .246 \quad x_2 \simeq .01 \quad y_2 = .240 = -z_3$$

Stage 3) $\quad z_{F_3} = .240 \quad f = 1/2 \quad \text{Slope} = -1$

$$\text{Intercept} = \frac{.240}{1/2} = .480 \quad x_3 \simeq .022 \quad y_3 = .461$$

2.F3. *Bubble Pt.* At P = 250 kPa. Want $\sum K_1 z_1 = 1$ .Solution uses DePriester chart for K values.

Guess $\quad T = -18^\circ C, \quad K_1 = 1, \quad K_2 = .043, \quad K_3 = .00095, \quad \sum = .52$

Converge to $\quad T = 0^\circ C$

Dew Pt. Calc. Want $\quad \sum \frac{z_1}{K_1} = 1.0$

Try $\quad T = 0^\circ C, \quad K_1 = 1.93, \quad K_2 = 0.11, \quad K_3 = 0.0033, \quad \sum = 120.26$

Converge to $\quad T = 124^\circ C$. This is a wide boiling feed.
T_{drum} must be lower than 95°C since that is feed temperature.

First Trial: Guess $T_{d,1} = 70^\circ C$: $K_1 = 7.8,\ K_2 = 1.07,\ K_3 = .083$

Guess $V/F = 0.5$. Rachford Rice Eq.

$$f\,V/F = \frac{(7.8-1)(.517)}{1+(6.8)(.5)} + \frac{(.07)(.091)}{1+(.07)(.5)} + \frac{(.083-1)(.392)}{1+(.083-1)(.5)} = .14$$

$$V/F = .6 \text{ gives } f(.6) = -.101$$

By linear interpolation: $V/F = .56.\ f(0.56) = -.0016$ which is close enough for first trial.

$$V = (V/F)F = 56, \quad L = 44$$

$$x_i = \frac{z_i}{1+(K_i-1)V/F} \text{ and } y_i = K_i x_i$$

$$x_1 = .1075 \quad x_2 = .088 \quad x_3 = .806 \quad \sum x = 1.001$$

$$y_1 = .839 \quad y_2 = .094 \quad y_3 = .067 \quad \sum y = .9999$$

Data: Pick $T_{ref} = 25^\circ C$. (Perry's 6th ed; p. 3-127), and (Perry's 6th ed; p. 3-138)

$$\lambda_1 = 81.76 \text{ cal/g } \times 44 = 3597.44 \text{ kcal/kmol}$$

$$\lambda_2 = 87.54 \text{ cal/g } \times 72 = 6302.88 \text{ kcal/kmol}$$

$$\lambda_3 = 86.80 \text{ cal/g } \times 114 = 9895.2 \text{ kcal/kmol}$$

at $T = 0^\circ C,\ C_{pL1} = 0.576 \text{ cal}/(\text{g } ^\circ C) \times 44 = 25.34 \text{ kcal/(kmol } ^\circ C)$.

For $T = 20 \text{ to } 123^\circ C,\ C_{pL3} = 65.89 \text{ kcal/(kmol } ^\circ C)$

at $T = 75^\circ C,\ C_{pL2} = 39.66 \text{ kcal/(kmol } ^\circ C)$. (Himmelblau/Appendix E-7)

$$C_{pv} = a + bT + cT^2$$

propane	a = 16.26	b = 5.398 × 10^{-2}	c = -3.134 × 10^{-5}
n-pentane	a = 27.45	b = 8.148 × 10^{-2}	c = -4.538 × 10^{-5}
**n-octane	a = 8.163	b = 140.217 × 10^{-3}	c = -44.127 × 10^{-6}

** Smith & Van Ness p. 106

Energy Balance: $E(T_d) = VH_v + Lh_L - Fh_F = 0$

$$Fh_F = 100\left[(.577)(25.34)+(.091)(39.66)+.392(65.89)\right](95.25) = 297,773 \text{ kcal/h}$$

$$Lh_L = 44\left[(.1075)(25.34)+(.088)(39.66)+(.806)(65.89)\right](70.25) = 117,450$$

$$VH_v = 56\Big[(.839)\left[3597.4+16.26+5.398\times10^{-2}(45)\right]$$
$$+(0.94)\left[6302.88+27.45+8.148\times10^{-2}(45)\right]$$
$$+(0.67)\left(9895.3+8.163+140.217\times10^{-3}(45)\right)\Big] = 240,423$$

$E(T_{drum}) = -60,101$ Thus, T_{drum} is too high.

Converge on $T_{drum} = 57.2^\circ C$: $K_1 = 6.4,\ K_2 = .8,\ K_3 = .054$

For $V/F = 0.513,\ f(0.513) = -0.0027.\ V = 51.3,\ L = 48.7$

$$x_1 = .137,\ x_2 = .101,\ x_3 = .762,\ \sum x_1 = 1.0000$$

$$y_1 = .878,\ y_2 = .081,\ y_3 = .041,\ \sum y_1 = 1.0000$$

$$Fh_F = 297,773;\ Lh_L = 90,459;\ \ VH_v = 209,999;\ \ E(T_{drum}) = +2685$$

Thus T_{drum} must be very close to 57.3°C.

$$x_1 = .136,\ \ x_2 = .101,\ \ x_3 = .762,\ y_1 = .328,\ \ y_2 = .081,\ \ y_3 = .041$$

$$V = 51.3 \text{ kmol/h, } L = 48.7 \text{ kmol/h}$$

Note: With different data T_{drum} may vary significantly.

2.F4. *New Problem 4th edition.* This is a mass and energy balance problem disguised as a flash distillation problem. Data is readily available in steam tables.. At 5000 kPa and 500K the feed is a liquid, h_F = 17.604 kJ/mol. For an adiabatic flash, $h_F = [VH_V + Lh_L]/F$

Vapor and liquid are in equilibrium. Saturated steam at 100 kPa is at T = 372.76K, h_L = 7.5214 kJ/mol, H_V = 48.19 kJ/mol

Mass balance: F = V + L where F in kmol/min = (1500 kg/min)(1 kmol/18.016 kg) = 83.259 kmol/min

EB: $Fh_F = VH_V + Lh_L \rightarrow$

(83.259 kmol/min)(17.604 kJ/mol)(1000 mol/kmol) = (48.19)(1000)V + (7.5214)(1000)L.

Solve equations simultaneously. L = 62.617 kmol/min = 1128.12 kg/min and V = 20.642 kmol/min = 371.88 kg/min

2.G1. Used Peng-Robinson for hydrocarbons.

Find $T_{drum} = 33.13^\circ C,\ L = 34.82$ and $V = 65.18$ kmol/h

In order ethylene, ethane, propane, propylene, n-butane, x_i (y_i) are:

$$0.0122(0.0748),\ 0.0866(0.3005),\ 0.3318(0.3781),\ 0.0306(0.0404),\ 0.5388(0.2062.)$$

2.G2. *New problem in 4th edition.* Part a. p = 31.26 kPa with $V/F)_{feed}$ = 0.0009903.

Part b. Use p_{feed} = 31.76 kPa, $V/F)_{feed}$ = 0.0

Part c. Drum p = 3.9 bar, T_{drmu} = 19.339, V/F = 0.18605,

Liquid mole fractions: C1 = 0.14663, C2 = 0.027869 (∑ = 0.05253 is in spec), C5 = 0.6171, C6 = 0.3404.

Vapor mole fractions: C1 = 0.68836, C2 = 0.20057, C5 = 0.9523, and C6 = 0.01584.

2.G3. *New problem 4th edition.* K values in Aspen Plus are higher by 17.6% (methane), 7.04% (n-butane) and 0.07% n-pentane. Since the K values are higher V/F is higher by 10.2%.

Results:

	x	y	K
Methane	0.004599	0.27039	58.79
n-butane	0.44567	0.52474	1.1774
n-pentane	0.54973	0.20488	0.37269

$V/F)_{drum}$ = 0.43419; $V/F)_{feed}$ = 0.3654; Q = -3183.4 cal/s

2.G4.

COMP	x(I)	y(I)
METHANE	0.12053E-01	0.84824
BUTANE	0.12978	0.78744E-01
PENTANE	0.29304	0.47918E-01
HEXANE	0.56513	0.25101E-01
V/F = 0.58354		

2.G5. *N.* Used NRTL. T = 368.07, Q = 14889 kW, 1st liquid/total liquid = 0.4221,

Comp	Liquid 1, x_1	Liquid 2, x_2	Vapor, y

Furfural	0.630	0.0226	0.0815
Water	0.346	0.965	0.820
Ethanol	0.0241	0.0125	0.0989

2.G6. Used Peng Robinson. Feed pressure = 10.6216 atm, Feed temperature = 81.14°C, V/F = 0.40001, Q_{drum} =0. There are very small differences in feed temperature with different versions of AspenPlus.

COMP	x(I)	y(I)
METHANE	0.000273	0.04959
BUTANE	0.18015	0.47976
PENTANE	0.51681	0.39979
HEXANE	0.30276	0.07086
V/F = 0.40001		

2.H1. *New Problem 4th ed* A. 563.4 R, b.V/F = .4066. c. 18.264 psia

2.H3. *New Problem. 4th ed.* Answer V/F = 0.564; xE = 0.00853, xhex = 0.421, x hept = .570; yE = .421, y Hex = 0.378, y Hept = .201.

2H4. *New Problem, 4th ed.* Answer: p_{drum} = 120.01, kPa = 17.40 psia
xB = 0.1561, xpen = 0.4255, x hept = 0.4184, yB = 0.5130, yPen = 0.4326 , yhept = 0.0544

2H5. *New problem 4th ed.*

a. SOLUTION. P = 198.52 kPa.
b. V/F = 0.24836, ethane x = 0.00337, y = 0.0824; Propane x =0.05069, y = 0.3539; Butane x= 0.1945, y = 0.3536; Pentane x = 0.3295, y = 0.1584; Hexane x = 0.3198, y = 0.0469 Heptane x = 0.1022, y = 0.00464
c. T = 34.48°C
d. T = -1.586°C and V/F = 0.0567

2H6. New problem in 4th ed.

C2.H6

	A	B	C	D	E	F	G	H	I	J
1										
2	Example 2-2 on spreadsheet									
3	K1	7	K2	2.4	K3	0.8	K4	0.3		
4	z1	0.3	z2	0.1	z3	0.15	z4	0.45		
5	Guess V/F	0.500823								
6	x1	0.074908	x2	0.058784	x3	0.166697	x4	0.692922		
7	y1	0.524353	y2	0.141081	y3	0.12	y4	0.207877	sum ↓	
8	yk-xi	0.449445		0.082297		-0.0467		-0.48505	-2.1E-07	
9								chk	-0.00021	
10		Goal seek 9I to zero by changing B5								

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