5. Cannibinol is only slightly soluble in methyl alcohol because the large hydrocarbon component of cannibinol negates the fact that they belong to the same family.

# Experiment 2

#### **CRYSTALLIZATION**

TIME ESTIMATE: Parts A and B (3 hours), Parts C (about 1hour)

CHEMICALS AND SUPPLIES PER 10 STUDENTS:

#### Part A

Impure sulfanilamide (5% fluorenone as the impurity) 10 g Grind thoroughly to make homogeneous.

95% Ethyl alcohol 250 mL

Filter paper for Büchner funnel

Melting point capillary tubes

Waste container for non-halogenated organic wastes.

#### Part B

The appropriate solvent for crystallizing the impure fluorene is methyl alcohol. Fluorene is too soluble in toluene and insoluble in water at all temperatures.

Impure fluorene (5% fluorenone as the impurity) 10 g Grind thoroughly to make homogeneous.

Methyl alcohol 300 mL

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Toluene 25 mL

Waste container for non-halogenated organic wastes.

## Part C

Acetylsalicylic acid	5 g
Benzoic acid	5 g
Benzoin	5 g
Dibenzoyl ethylene	5 g
Succinimide	5 g
o-Toluic acid	5 g

Prepare unknowns consisting of pure samples of the above 6 compounds; about 0.2 g for each unknown sample.

# Part D (Answers)

1. Phenanthrene

95% ethyl alcohol - best solvent water - not soluble toluene - too soluble

2. Cholesterol

ether - too soluble 95% ethyl alcohol - best solvent water - not soluble

3. Acetaminophen

95% ethyl alcohol - too soluble water - best solvent toluene - not very soluble

4. Urea

Water - too soluble 95% ethyl alcohol - best solvent Hexane - not very soluble

## **CAS Registry numbers:**

Sulfanilamide 63-74-1

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Acetanilide	103-84-4
95% Ethyl alcohol (Ethanol)	64-17-5
Fluorene	86-73-7
Fluorenone	486-25-9
Methyl alcohol (methanol)	67-56-1
Toluene	108-88-3
Acetylsalicylic acid	50-78-2
Benzoic acid	65-85-0
Benzoin	119-53-9
Dibenzoyl ethylene	4070-75-1
Succinimide	123-56-8
o-toluic acid	118-90-1

## SPECIAL NOTES

In the Pre-lab Calculations for Part A, students calculate the amount of sulfanilamide which will remain in the mother liquor. If they perform the Optional Exercise in Part A, they determine the weight of solid in the mother liquor. However, the actual weight of solid in the mother liquor is usually much greater than the amount calculated in the Pre-lab Calculations. This is because the calculation does not take into account the impurity, which ends up in the mother liquor. Also, the calculation assumes that a minimum amount of solvent is used to dissolve the impure sulfanilamide at 78 °C. It is likely that most students use more than the minimum amount.

# ANSWERS TO QUESTIONS

- 1. Too much solvent was added. Since 10 mL of 95% ethyl alcohol will dissolve 0.14 g of sulfanilamide at 0 °C, none of the 0.1 g of sulfanilamide will crystallize when the solution is cooled. To make the crystallization work, the excess solvent must be evaporated.
- 2. The boiling point of the solvent is higher than the melting point of fluorenol. While performing this crystallization, it is possible that the fluorenol would melt rather than dissolve, thus forming an oil which could be difficult to crystallize.
- 3. Biphenyl is highly soluble in both hot and cold benzene. The solubility curve would like **C** in Figure 11.1 on page 665 of the Textbook.

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